

Ecology Unit Exam Study Guide

Topics include: Energy Flow, food webs, energy pyramids, Nutrient Cycles, Ecological relationships such as parasitism and mutualism, Biodiversity, Anything related to the biodiversity labs, Competitive Exclusion principle, habitat and niche, Carrying Capacity and human population growth, Ecological succession, Algae blooms, dead zones and limiting nutrient, Biomes, Invasive species, and Pollution

Energy Flow + food webs + energy pyramids

- Sun is the main source of energy
- Autotrophs (producers)
 - Make own energy/food (not from other organisms)
 - Photosynthesis: Energy from sunlight
 - Chemosynthesis: energy made during chemical reactions
 - EX: Trees, Plants, Algae
- Decomposers: break down organic matter such as dead organism remains
 - Allows other organisms to use these materials from decaying materials for own use.
- Heterotrophs (consumers): need to eat other organisms in order to get the materials you need to maintain homeostasis
 - Primary consumer (1)- eats the producer
 - secondary consumers (2)- eats 1 consumer
 - Tertiary consumer- the TOP consumer that eats secondary consumer

Energy Pyramid

- Represents the available amount of energy at each trophic level
- Energy moves in ONE direction
- Most energy at the bottom of the pyramid
- 10% Rule
 - 90% of energy is used for bodily functions (ex. Digestion and movement) and is released as heat
 - Only 10% is transferred to the next level. Energy is lost as you move up.
 - When you consume food, energy turns into HEAT.

Unit 2 Study Guide Ecology

Mr. Rohit Manglik



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Introductory Ecology Mr. Rohit Manglik, 2024-04-06 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels **Resources in Education**, 2000 **X-kit FET Grade 11 BUSINESS STUDIES**, 2007

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Field and Laboratory Methods for General Ecology James E Brower, Jerrold H Zar, Carl N. von Ende, 1998 This introductory ecology lab manual focuses on the process of collecting recording and analyzing data and equips students with the tools they need to function in more advanced science courses It reflects the most current techniques for data gathering so that students can obtain the most accurate samples Balanced coverage of plant animal and physical elements offers a diverse range of exercises Includes exercise on writing research reports

Biosafety in the Laboratory Division on Engineering and Physical Sciences, Commission on Physical Sciences, Mathematics, and Applications, Committee on Hazardous Biological Substances in the Laboratory, National Research Council, 1989-01-01 Biosafety in the Laboratory is a concise set of practical guidelines for handling and disposing of biohazardous material The consensus of top experts in laboratory safety this volume provides the information needed for immediate improvement of safety practices It discusses high and low risk biological agents including the highest risk materials handled in labs today presents the seven basic rules of biosafety addresses special issues such as the shipping of dangerous materials covers waste disposal in detail offers a checklist for administering laboratory safety and more

Monthly Catalogue, United States Public Documents ,1994

[McGraw-Hill encyclopedia of science & technology](#) McGraw-Hill, 2002

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ/mole}$

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multiply the mass by the change in temperature and the c given trial 1 use Hess's law and the following equations and ΔH values to determine the heat of reaction for the reaction c_2h_4

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web thus for reaction 9.2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove Hess's law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 Hess's law in this experiment you will measure the amount of heat released in these three related exothermic reactions $1. NaOH(aq) + HCl(aq) \rightarrow NaCl(aq) + H_2O(l)$ $2. NaOH(aq) + H_2SO_4(aq) \rightarrow Na_2SO_4(aq) + H_2O(l)$ $3. HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$

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