

charge on the carbon makes it polar.

Electrophile is an **electron-pair donor**.  
Electrophile is an electron pair to somewhere without enough electrons.  
Electrophiles are **electron pair acceptors**.  
 $\text{H}^+$  and  $\text{NH}_3$  are all **nucleophiles** that can react with haloalkanes.



## Haloalkanes Can Undergo Nucleophilic Substitution Reactions

Haloalkanes can react with a polar molecule by kicking out the functional group and taking its place.  
This is called a **nucleophilic substitution reaction**. It works like this:



X stands for one of the halogens (F, Cl, Br or I).  
Nu stands for a nucleophile.

The lone pair of electrons on the nucleophile attacks the  $\delta^+$  carbon.

The C-X bond breaks.

The halogen leaves, taking both electrons with it. A new bond forms between the carbon and the nucleophile.

Product of these reactions depends on what the nucleophile is...

## Haloalkanes React with Hydroxides to Form Alcohols

Haloalkanes, bromoethane can react to form ethanol. You have to use warm sodium or potassium hydroxide — it's a nucleophilic substitution reaction.

This is sometimes called **hydrolysis**, because exactly the same will happen with other haloalkanes (Hydrolysis is splitting a molecule).

# Unit 3 Section 1 Quiz Chemistry

**Deborah Herridge**



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