

## Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:  $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before  $\underline{\quad 2 \quad}$   $\underline{\text{ex}}$   $\underline{\quad 0 \quad}$   $\underline{\quad 0 \quad}$  3.0 mol O<sub>2</sub>

Change  $\underline{\quad -2 \quad}$   $\underline{\quad -3 \quad}$   $\underline{\quad 2 \quad}$   $\underline{\quad 2 \quad}$

After  $\underline{\quad 0 \quad}$   $\underline{\text{ex}}$   $\underline{\quad 2 \quad}$   $\underline{\quad 2 \quad}$

- Propane, C<sub>3</sub>H<sub>8</sub>, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation:  $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5  $\underline{\text{ex}}$  0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0  $\underline{\text{ex}}$  0.5 0.75

- Ammonia, NH<sub>3</sub>, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation:  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before  $\underline{\text{ex}}$  1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After  $\underline{\text{ex}}$  0 2.5

# Unit 7 Mole Relationships

**A Loxley**



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## **Table of Contents Unit 7 Mole Relationships**

1. Understanding the eBook Unit 7 Mole Relationships
  - The Rise of Digital Reading Unit 7 Mole Relationships
  - Advantages of eBooks Over Traditional Books
2. Identifying Unit 7 Mole Relationships
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 7 Mole Relationships
  - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 7 Mole Relationships
  - Personalized Recommendations
  - Unit 7 Mole Relationships User Reviews and Ratings
  - Unit 7 Mole Relationships and Bestseller Lists
5. Accessing Unit 7 Mole Relationships Free and Paid eBooks

- Unit 7 Mole Relationships Public Domain eBooks
- Unit 7 Mole Relationships eBook Subscription Services
- Unit 7 Mole Relationships Budget-Friendly Options
- 6. Navigating Unit 7 Mole Relationships eBook Formats
  - ePub, PDF, MOBI, and More
  - Unit 7 Mole Relationships Compatibility with Devices
  - Unit 7 Mole Relationships Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Unit 7 Mole Relationships
  - Highlighting and Note-Taking Unit 7 Mole Relationships
  - Interactive Elements Unit 7 Mole Relationships
- 8. Staying Engaged with Unit 7 Mole Relationships
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Unit 7 Mole Relationships
- 9. Balancing eBooks and Physical Books Unit 7 Mole Relationships
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit 7 Mole Relationships
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 7 Mole Relationships
  - Setting Reading Goals Unit 7 Mole Relationships
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 7 Mole Relationships
  - Fact-Checking eBook Content of Unit 7 Mole Relationships
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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