

Reaction Quotient

How do you predict which direction a reaction will proceed to reach equilibrium?

Why?

When a reaction reaches equilibrium there must be some non-negligible amount of every species in the reaction, otherwise the reaction cannot react in both directions. Knowing this, it is very easy to predict which direction a reaction will go to reach equilibrium when one of the components of the reaction has an initial concentration of zero. Many of the problems you have worked with thus far have some component at zero concentration, but real life does not work that way. Most of the time, the reaction in question has some measurable amount of every species. Deciding which way the reaction will go to reach equilibrium then becomes more challenging.

Model 1 – A Theoretical Equilibrium

Trial 1	A (g) + B (g) \leftrightarrow C (g)		
Initial	1.000 M	1.000 M	1.000 M
Change	\leftarrow		
Equilibrium	1.464 M	1.464 M	0.536 M

Trial 3	A (g) + B (g) \leftrightarrow C (g)		
Initial	1.000 M	0.500 M	1.500 M
Change			
Equilibrium	1.864 M	1.364 M	0.636 M

Trial 5	A (g) + B (g) \leftrightarrow C (g)		
Initial	1.400 M	1.200 M	0.400 M
Change			
Equilibrium	1.388 M	1.188 M	0.412 M

Trial 2	A (g) + B (g) \leftrightarrow C (g)		
Initial	2.000 M	0.500 M	0.500 M
Change			
Equilibrium	2.150 M	0.650 M	0.350 M

Trial 4	A (g) + B (g) \leftrightarrow C (g)		
Initial	1.600 M	1.000 M	0.400 M
Change			
Equilibrium	1.600 M	1.000 M	0.400 M

Trial 6	A (g) + B (g) \leftrightarrow C (g)		
Initial	0.750 M	2.000 M	0.250 M
Change			
Equilibrium	0.675 M	1.925 M	0.325 M

- Examine Model 1.
 - Write the theoretical chemical reaction that is used in the trials of Model 1.
 - If 0.50 M of reactant A reacts, predict the change in concentration of B and C.
- What variables were changed in the different trials shown in Model 1?

Pogil Activities For High School Chemistry Equilibrium

AN Whitehead



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Table of Contents Pogil Activities For High School Chemistry Equilibrium

1. Understanding the eBook Pogil Activities For High School Chemistry Equilibrium
 - The Rise of Digital Reading Pogil Activities For High School Chemistry Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For High School Chemistry Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For High School Chemistry Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For High School Chemistry Equilibrium

- Personalized Recommendations
 - Pogil Activities For High School Chemistry Equilibrium User Reviews and Ratings
 - Pogil Activities For High School Chemistry Equilibrium and Bestseller Lists
5. Accessing Pogil Activities For High School Chemistry Equilibrium Free and Paid eBooks
 - Pogil Activities For High School Chemistry Equilibrium Public Domain eBooks
 - Pogil Activities For High School Chemistry Equilibrium eBook Subscription Services
 - Pogil Activities For High School Chemistry Equilibrium Budget-Friendly Options
 6. Navigating Pogil Activities For High School Chemistry Equilibrium eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activities For High School Chemistry Equilibrium Compatibility with Devices
 - Pogil Activities For High School Chemistry Equilibrium Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For High School Chemistry Equilibrium
 - Highlighting and Note-Taking Pogil Activities For High School Chemistry Equilibrium
 - Interactive Elements Pogil Activities For High School Chemistry Equilibrium
 8. Staying Engaged with Pogil Activities For High School Chemistry Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For High School Chemistry Equilibrium
 9. Balancing eBooks and Physical Books Pogil Activities For High School Chemistry Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For High School Chemistry Equilibrium
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Activities For High School Chemistry Equilibrium
 - Setting Reading Goals Pogil Activities For High School Chemistry Equilibrium
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Activities For High School Chemistry Equilibrium

- Fact-Checking eBook Content of Pogil Activities For High School Chemistry Equilibrium
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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