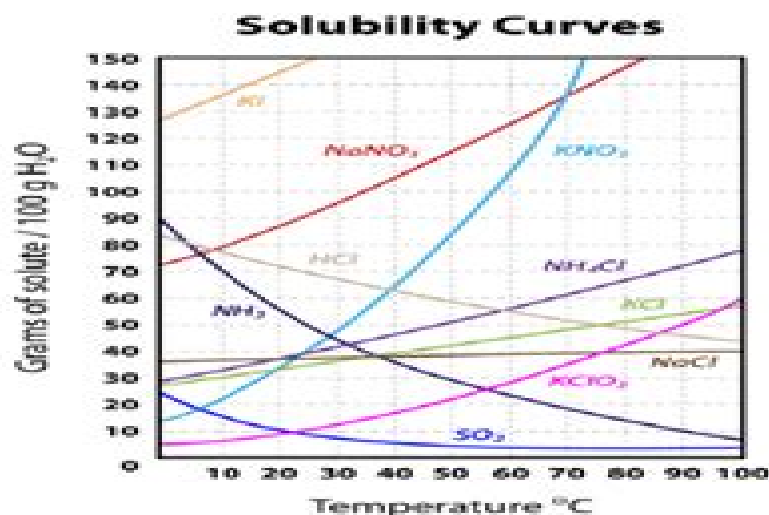


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 ml.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO_3 at 10°C , 40°C , & 80°C ?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH_3 can you dissolve in water at 10°C , 30°C , & 90°C ?
@ 10°C = ~70 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C , 50°C , & 90°C ?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C ?
About 40 grams
- Except for KCl What substance is the most soluble in water at 20°C ?
 NaNO_3
- Except for KCl what substance is the most soluble in water at 72°C ?
 KNO_3
- Look up The boiling points for SO_2 , HCl , and NH_3 and write them below:
B.P. SO_2 = -10°C
B.P. HCl = -85°C
B.P. NH_3 = -33°C
- Based on the boiling points you looked up, are SO_2 , HCl , & NH_3 solids, liquids, or gases at 0°C ? (and anywhere else on the solubility curve)
gases

Pogil Interpreting Solubility Curves Answer Key

**Juliette Lantz, Renée Cole, The POGIL
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Table of Contents Pogil Interpreting Solubility Curves Answer Key

1. Understanding the eBook Pogil Interpreting Solubility Curves Answer Key
 - The Rise of Digital Reading Pogil Interpreting Solubility Curves Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Interpreting Solubility Curves Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Interpreting Solubility Curves Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Interpreting Solubility Curves Answer Key
 - Personalized Recommendations
 - Pogil Interpreting Solubility Curves Answer Key User Reviews and Ratings
 - Pogil Interpreting Solubility Curves Answer Key and Bestseller Lists
5. Accessing Pogil Interpreting Solubility Curves Answer Key Free and Paid eBooks
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 - Pogil Interpreting Solubility Curves Answer Key eBook Subscription Services

- Pogil Interpreting Solubility Curves Answer Key Budget-Friendly Options
- 6. Navigating Pogil Interpreting Solubility Curves Answer Key eBook Formats
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 - Pogil Interpreting Solubility Curves Answer Key Compatibility with Devices
 - Pogil Interpreting Solubility Curves Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Interpreting Solubility Curves Answer Key
 - Highlighting and Note-Taking Pogil Interpreting Solubility Curves Answer Key
 - Interactive Elements Pogil Interpreting Solubility Curves Answer Key
- 8. Staying Engaged with Pogil Interpreting Solubility Curves Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Interpreting Solubility Curves Answer Key
- 9. Balancing eBooks and Physical Books Pogil Interpreting Solubility Curves Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Interpreting Solubility Curves Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Interpreting Solubility Curves Answer Key
 - Setting Reading Goals Pogil Interpreting Solubility Curves Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Interpreting Solubility Curves Answer Key
 - Fact-Checking eBook Content of Pogil Interpreting Solubility Curves Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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