

Module Overview

What are the coefficients in a chemical equation for combustion?

What?

A chemical equation represents what we see when we measure reactants and products produced from a reaction. Coefficients in formulas indicate there are thousands of molecules being produced or consumed. Typically, when we balance the chemical equation we don't see measured quantities produced. However, we need the stoichiometry represented by coefficients because we can't measure, measure all those. Most of what you see (large numbers) are the biggest numbers of each chemical in the equation. You can see how many chemical equations are there that make coefficients represent what is happening in the world. In this activity you will explore how different types of chemical reactions can be represented.

Activity 1 – A Chemical Reaction



1. Consider the reaction in Model 1.

a. What are the coefficients for each of the following substances in the reaction?



b. Does your coefficient value in balance the equation in Model 1?

2. Consider each reaction below as a chemical reaction in Model 1.

a. Halving the amount of reaction consumed and produced water.

b. Doubling the rate of H_2 or O_2 or H_2O . Balance the equation the fastest while maintaining quality.

	H_2 Coefficient	O_2 Coefficient	H_2O Coefficient	Ratio $H_2 : O_2 : H_2O$ (simplified)
How a single product from many reactants? (single substance would be consumed or produced)				
If the reaction consumed less than double water, how many molecules would be consumed or produced?				
If the reaction consumed 1/100th more than double water, how many molecules would be consumed or produced?				

Pogil Mole Ratios Worksheet

Sunil Tanna

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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