

Phase Changes in Pure Substances

WHY?

A phase diagram is a convenient way to represent which phase (solid, liquid, or gas) or combination of phases is present at different temperatures and pressures. Knowing the relationship of temperature and pressure to a material's phase is necessary in order to determine its suitability for a particular application (e.g., a high-temperature engine component or a cooking oil). Phase diagrams are used, for example, to plan the synthesis of artificial diamonds and new materials, or to deduce the history of geological samples. Additionally, knowledge of a material's capacity to absorb energy in a particular phase or while changing phases is essential to the temperature and stability of systems involving chemical reactions and energy transfer.

LEARNING OBJECTIVES

- Interpret the major features of a phase diagram
- Interpret a heating or cooling curve to determine the physical and thermochemical properties of a pure substance

SUCCESS CRITERIA

- Identify stable phases, at particular temperatures and pressures, from a phase diagram
- Use a phase diagram to identify the temperatures and pressures at which phase changes occur
- Use a heating curve to determine a substance's melting and boiling points, specific heat capacities for each phase, and enthalpies associated with phase changes

PREREQUISITES

06-1 *Thermochemistry and Calorimetry*

06-2 *Internal Energy and Enthalpy*

INFORMATION



A *phase diagram* summarizes the relationships between a substance's phase, temperature, and pressure. It represents a map of the states of matter for a pure substance. Within a pure phase, the addition or removal of heat results in a change in temperature.

Boundary lines between pure phases represent the temperatures and pressures at which two phases are in equilibrium. As a phase change or transition occurs across these boundary lines, the temperature does not change until the phase change is complete.

A *triple point* marks the temperature and pressure where three boundary lines come together. At these points, three phases are in equilibrium.

A phase diagram can be produced by observing the which phase of a substance is present under specific temperature and pressure conditions, as heat energy is added to or removed from the system. A *heating curve* is a plot of the temperature of the substance as a function of the energy added. A *cooling curve* is a plot of the temperature as a function of the energy removed.

Measurement of the amount of energy transferred in a heating or cooling experiment makes it possible to determine the enthalpies of fusion and vaporization for phase changes, as well as the specific heat capacities of different pure phases.

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