

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

*On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.*

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles $\text{KClO}_3$ needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles $\text{HCl}$ needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200	+0.100	+0.100	
After:	0	?	0.100	0.100	

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	-7.70	+4.40	+6.60
After:	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	-0.0040	-0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# Review Sheet Unit Answers Chemistry Answers

**James E. Brady, Fred Senese**



## Review Sheet Unit Answers Chemistry Answers:

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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*Division Bulletin* United States. Division of Vocational Education,1939 **Study Guide to Accompany Calculus for the Management, Life, and Social Sciences** Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences *Resources in Education* ,1991-04 Vocational Division Bulletin United States. Division of Vocational and Technical Education,1939 Vocational Education Bulletin United States. Division of Vocational

Education,1961 *Curriculum Materials for Trade and Industrial Education, 1963* ,1964 **Study Guide to Accompany Basics for Chemistry** Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to

particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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