Stoichiometry Review:

Test Review



	the discovery chamical
1.	What is stoichiomorry? [] Study of quantitative relationships in chamical
2.	On which law is stoichiometry based? [] Law of conserv action of mores.
3.	According to the law of conservation of mass, the mass of the must must
4.	What is a mole ratio? [] fractions/rectio comparing two substances.
5.	Given the following balanced equation: 2Al + 3Br ₂ → 2AlBr ₁ , determine the mole ratio used for
	calculating moles of aluminum bromide from moles of bromine. []2 rool Al Br /3 rool BQ
6.	What must you have before beginning any stoichiometric problem? [] balanced equation
7.	What is a limiting reactant? [] Yunk out
8	What is an excess reactant? [] Left over
9.	A booklet is prepared from 2 covers, 3 staples, and 20 pages. Assume we have 60 covers, 120
	staples, and 400 pages.
	a. What is the largest number of complete booklets that can be made? [20]
	b. Which is the "limiting reactant" in this system? [copy]
10	How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction? [2] LiBr + [1] Cl ₂ → [2] LiCl + [1] Br ₃
O	anol Lier = 0.046 mol Lich
	A mol Li Br Hydrogen peroxide breaks down, releasing oxygen, in the following reaction. What mass of oxygen is produced when 1.840 mol of H ₂ O ₂ decomposes? [2] 1H ₂ O ₂ → [3] 1H ₂ O + [1] 1O ₂
8	340 molth202 1 mol Od 132.09 08 = 24.449 02
	la mal H O. I mal O.

 Hydrogen is generated by passing hot steam over iron, which oxidizes to form Fe₂O₆, in the following equation. If 625 g of Fe₂O₆ is produced in the reaction, how many moles of hydrogen are produced at the same time?

Review Stoichiometry Section 1 Answer Key

Annelies Wilder-Smith

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