

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation: $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before $\underline{\quad 2 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5 $\underline{\quad \text{ex} \quad}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\quad \text{ex} \quad}$ 0.5 0.75

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before $\underline{\quad \text{ex} \quad}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\quad \text{ex} \quad}$ 0 2.5

Unit 7 Mole Relationships Answer Key

Ralph Thomas Myers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Cehmistry Textbook for College and University USA* Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA

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Table of Contents Unit 7 Mole Relationships Answer Key

1. Understanding the eBook Unit 7 Mole Relationships Answer Key
 - The Rise of Digital Reading Unit 7 Mole Relationships Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 7 Mole Relationships Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 7 Mole Relationships Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 7 Mole Relationships Answer Key
 - Personalized Recommendations
 - Unit 7 Mole Relationships Answer Key User Reviews and Ratings
 - Unit 7 Mole Relationships Answer Key and Bestseller Lists
5. Accessing Unit 7 Mole Relationships Answer Key Free and Paid eBooks
 - Unit 7 Mole Relationships Answer Key Public Domain eBooks
 - Unit 7 Mole Relationships Answer Key eBook Subscription Services
 - Unit 7 Mole Relationships Answer Key Budget-Friendly Options
6. Navigating Unit 7 Mole Relationships Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 7 Mole Relationships Answer Key Compatibility with Devices
 - Unit 7 Mole Relationships Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 7 Mole Relationships Answer Key
 - Highlighting and Note-Taking Unit 7 Mole Relationships Answer Key
 - Interactive Elements Unit 7 Mole Relationships Answer Key
8. Staying Engaged with Unit 7 Mole Relationships Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Unit 7 Mole Relationships Answer Key
- 9. Balancing eBooks and Physical Books Unit 7 Mole Relationships Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 7 Mole Relationships Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 7 Mole Relationships Answer Key
 - Setting Reading Goals Unit 7 Mole Relationships Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 7 Mole Relationships Answer Key
 - Fact-Checking eBook Content of Unit 7 Mole Relationships Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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