

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation: $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before $\underline{\quad 2 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\text{ex}}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5 $\underline{\text{ex}}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\text{ex}}$ 0.5 0.75

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before $\underline{\text{ex}}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\text{ex}}$ 0 2.5

Unit 8 Ws Mole Relationships Answers

Karin Nielsen-Saines



Unit 8 Ws Mole Relationships Answers:

Selected Works of Paul J. Flory Volume I, Single Best Answer MCQs in Anaesthesia Cyprian Mendonca,,Mahesh Chaudhari,,A Pitchiah,2011-09-01 This book comprises six sets of single best answer practice papers Each set contains 30 single best answer questions on physiology pharmacology clinical measurement and physics The scenarios are based on the application of a wide knowledge of basic sciences relevant to the clinical practice of anaesthesia The best possible answer to a given question is substantiated by detailed explanation drawn from recent journal articles and textbooks of anaesthesia and basic sciences These questions enable the candidates to assess their knowledge in basic sciences and their ability to apply it to clinical practice Alongside the previously published book Single Best Answer MCQs in Anaesthesia Volume I Clinical Anaesthesia ISBN 978 1 903378 75 5 this book is an ideal companion for candidates sitting postgraduate examinations in anaesthesia intensive care medicine and pain management It will also be a valuable educational resource for all trainees and practising anaesthetists

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg,2015-01-16

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Alkaline Earth Hydroxides in Water and Aqueous Solutions I. Lambert,H.L. Clever,2013-10-22 This volume contains evaluated data on the solubility of beryllium hydroxide magnesium hydroxide calcium hydroxide strontium hydroxide and barium hydroxide in water and in a number of electrolyte and nonelectrolyte solutions in water The alkaline earth hydroxides can be divided into two groups depending on the hydration of the solid First the sparingly soluble anhydrous beryllium magnesium and calcium hydroxides whose freshly precipitated solids are poorly crystalline and show decreasing solubility with aging and whose solubility in water decreases with increasing temperature Second the soluble strontium and barium hydroxide octahydrates that form crystalline precipitates which do not show changes in solubility on aging and whose solubility in water increases with increasing temperature

Uranium · Plutonium Transplutonic Elements Harold C. Hodge,J.N. Stannard,J.B. Hursh,2013-06-29 More than a year ago the three editors sat down at a table and worked out a set of six chapter headings which they believed might serve in turn for each of the three sections of this handbook The reader will note a similarity in order of presentation and in emphasis However as our editorial plans progressed it became apparent that for each element and for the element group there were one or two special topics appropriate for that section alone Accordingly in the section on uranium the common pattern holds for Chaps 1 through 6 which include an introduction

Chap 1 a discussion of the physical and chemical properties Chap 2 experimental data on animals Chap 3 experimental data on man Chap 4 the rationale and development of air concentration limits to control industrial worker exposure Chap 5 and the practical problems of applying such limits in the uranium industry Chap 6 Chap 7 entitled Uranium Mining Hazards is the subject category which is special for uranium the chapter brings up to date the account of an important occupational hazard which was first noted by GEORGIUS AGRICOLA 1490 1555

Chemical Thermodynamics for Earth Scientists Philip Fletcher, 1993 A text providing a systematic introduction to the application of thermodynamics to chemical reactions occurring in the earth's geological environment for undergraduate and postgraduate geochemistry geology and environmental science students It covers three essential topics thermodynamic principles thermodynamic properties of geological materials and the use of thermodynamic data with graphical and computational techniques to predict properties of mineral and fluid assemblies Annotation copyright by Book News Inc Portland OR

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ENGINEERING CHEMISTRY FOR DIPLOMA RANJAN KUMAR MOHAPATRA, 2014-09-10 This book is written strictly for the first and second semester diploma students of engineering chemistry according to the revised syllabus It aims to provide a thorough understanding of the chemical concepts theories and principles in Engineering Chemistry in a clear and concise manner so that the average students are able to grasp the intricacies of the subject Explaining general concepts of atomic structure and chemical bond the book covers all advanced topics such as acid base theory concentration of solutions electrochemistry corrosion metallurgy hydrocarbons sources of water and its treatment lubricants and adhesives fuel polymer and environmental chemistry Each theoretical concept is well supported by illustrative examples Besides the book provides a large number of solved problems to reinforce the theoretical understanding of concepts Each chapter contains glossary terms and provides short questions and long questions for practice Previous year question papers and model questions with answers are appended at the end of the book to help students ace in examinations

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Adopting the Track of Appearance: An Emotional Symphony within **Unit 8 Ws Mole Relationships Answers**

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