

Study Guide – Chapter 3: Stoichiometry

Test Date: Thursday, 12/18/14

- **Chemical Equations**
 - Writing word equations
 - remember to balance & include the states of matter
 - Balancing chemical equations w/coefficients
 - 5 types of reactions
 - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
 - identify what type of reaction is happening from a given equation
 - predict the products when given the reactants or told the type of reaction...
 - $C_xH_y + O_2(g)$ combustion forms products of $CO_2(g)$ and $H_2O(g)$
 - Synthesis of metal + nonmetal forms an ionic compound (criss cross charges for formula)
 - Decomposition breaks compound into its elements (watch out for diatomic BrINCiHOFs or P_4 / S_8 solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
 - Calculations of formula weights and/or molar masses
 - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
 - Know the difference between an empirical and a molecular formula
 - Three types of problems
 - Normal problems will give % composition data or mass data to start
 - Combustion Analysis (like the quiz)
 - Hydrate Determination
 - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
 - Dimensional Analysis!!!
 - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec. CH_4)
 - use the molar mass to go from mass to moles and/or moles to mass
 - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
 - through calculations, identify which is the limiting reactant and which is the excess reactant
 - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you ALWAYS choose the LESSER quantity calculated)
 - determine how much of the excess reactant is used and how much is left over after the reaction is completed
 - calculate % yield
 - calculate theoretical yield and given the actual yield (from lab) of the substance
 - given % yield, you can determine actual yield collected in the lab

Unit 11 Stoichiometry Study Guide Answer Key

Marco Cascella



Unit 11 Stoichiometry Study Guide Answer Key:

Decoding **Unit 11 Stoichiometry Study Guide Answer Key**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Unit 11 Stoichiometry Study Guide Answer Key**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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