

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

| | | |
|------------------|------------------------|-------------------------------|
| energy | temperature | standard state |
| potential energy | heat | standard heat of formation |
| kinetic energy | specific heat capacity | exothermic |
| enthalpy | endothermic | Law of conservation of Energy |
| thermal | Hess's Law | isolated, closed, open system |

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$Q \text{ lost (Al)} = Q \text{ gained (Water)}$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

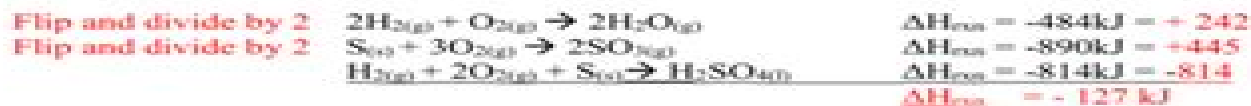
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

| | | |
|----------------------------------|--|---|
| $q = ?$ | $\Delta H = q/n$ | $n \text{ CH}_3\text{OH} = m/\text{MM}$ |
| $m = 300 \text{ g}$ | $q = H \times n$ | $n = 300 \text{ g} / 32 \text{ g/mol}$ |
| $\Delta H = 35.3 \text{ kJ/mol}$ | $q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$ | $n = 9.375 \text{ mol}$ |
| | $q = 332.81 \text{ kJ}$ | |

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



Unit 11 Test Thermochemistry

ML Morrison



Unit 11 Test Thermochemistry:

This book delves into Unit 11 Test Thermochemistry. Unit 11 Test Thermochemistry is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Unit 11 Test Thermochemistry, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Unit 11 Test Thermochemistry
- Chapter 2: Essential Elements of Unit 11 Test Thermochemistry
- Chapter 3: Unit 11 Test Thermochemistry in Everyday Life
- Chapter 4: Unit 11 Test Thermochemistry in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Unit 11 Test Thermochemistry. The first chapter will explore what Unit 11 Test Thermochemistry is, why Unit 11 Test Thermochemistry is vital, and how to effectively learn about Unit 11 Test Thermochemistry.

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6. In chapter 5, the author will draw a conclusion about Unit 11 Test Thermochemistry. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Unit 11 Test Thermochemistry.

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