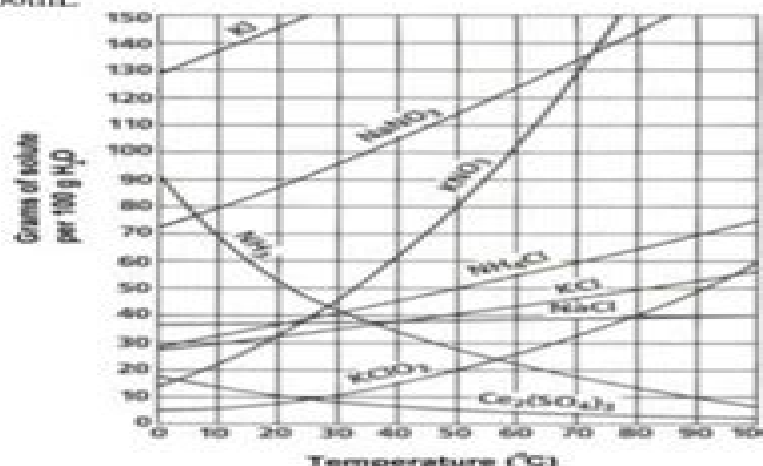


Name: _____ Date: _____ Period: _____

Solubility Curve Practice Problems

Directions: Find the mass of solute will dissolve in 100mL of water at the following temperatures?

1. KNO_3 at 70°C = _____
2. NaCl at 100°C = _____
3. NH_4Cl at 90°C = _____
4. Which of the **above** three substances is most soluble in water at 15°C . = _____
5. Which of the **above** three substances is least soluble in water at 15°C . = _____



2. Solution	Saturated, super saturated or Unsaturated?	If unsaturated: How much more solute can dissolve in the solution?
a solution that contains 70g of NaNO_3 at 30°C (in 100 mL H_2O)		
a solution that contains 50g of NH_4Cl at 50°C (in 100 mL H_2O)		
a solution that contains 20g of KClO_3 at 50°C (in 50 mL H_2O)		
a solution that contains 70g of KI at 0°C (in 200 mL H_2O)		

3. a. How many grams of KCl can dissolve at 5°C ? _____
 b. How many grams of KCl can dissolve at 25°C ? _____
 c. How many grams of $\text{Ce}_2(\text{SO}_4)_3$ can dissolve at 10°C ? _____
 d. How many grams of $\text{Ce}_2(\text{SO}_4)_3$ can dissolve at 50°C ? _____
4. At 90°C , you dissolved 10 g of KCl in 100. g of water. Is this solution saturated or unsaturated?
5. A mass of 100 g of NaNO_3 is dissolved in 100 g of water at 80°C .
 a) Is the solution saturated or unsaturated? _____
 b) As the solution is cooled, at what temperature should solid first appear in the solution? Explain.
6. Use the graph to answer the following two questions:
 Which compound is most soluble at 20°C ?
 Which is the least soluble at 40°C ?
 Which substance on the graph is **least** soluble at 10°C ?

Unit 12 Answers Solubility

Roger Slade, Maureen Slade



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