

A solution of HCl is to be prepared from 1.45 mol H^+ . What volume of solution can be produced using 9.87 g HCl?



$$9.87g HCl \times \frac{1 \text{ mol HCl}}{36} \times \frac{1 \text{ mol } H^+}{1 \text{ mol HCl}} \times \frac{1 L}{1.45 \text{ mol } H^+} \times \frac{1000 \text{ mL}}{1 L} = \boxed{189.08 \text{ mL}}$$

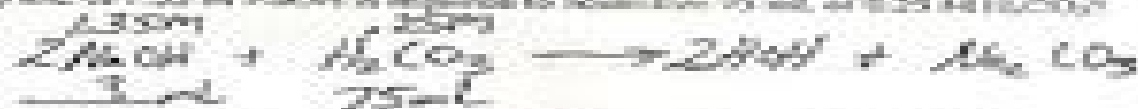
What mass of solid AgBr is produced when 30.0 mL of 0.15 M $AgNO_3$ is added to 10.0 mL of 0.10 M NaBr?



$$\text{Solid AgBr} = \frac{1}{100 \text{ mL}} \times \frac{1 L}{1000 \text{ mL}} \times \frac{0.15 \text{ mol AgNO}_3}{1 L} \times \frac{1 \text{ mol AgBr}}{1 \text{ mol AgNO}_3} \times \frac{188 \text{ g AgBr}}{1 \text{ mol}} = \boxed{1.41 \text{ g AgBr}}$$

$$\text{Solid AgBr} = \frac{1}{100 \text{ mL}} \times \frac{1 L}{1000 \text{ mL}} \times \frac{0.10 \text{ mol NaBr}}{1 L} \times \frac{1 \text{ mol AgBr}}{1 \text{ mol NaBr}} \times \frac{188 \text{ g AgBr}}{1 \text{ mol}} = 1.88 \text{ g AgBr}$$

How many mL of 1.35 M HNO_3 is required to neutralize 75 mL of 0.25 M Na_2CO_3 ?



$$75 \text{ mL } Na_2CO_3 \times \frac{1 L}{1000 \text{ mL}} \times \frac{0.25 \text{ mol } Na_2CO_3}{1 L} \times \frac{2 \text{ mol HNO}_3}{1 \text{ mol } Na_2CO_3} \times \frac{1 L}{1.35 \text{ mol HNO}_3} \times \frac{1000 \text{ mL}}{1 L} = \boxed{27.78 \text{ mL}}$$

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Clemens Wendtner



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