

3. When mercury (II) nitrate is heated, it decomposes to form mercury (II) oxide, nitrogen dioxide, and oxygen gas according to the following BALANCED equation.
- $$2 \text{Hg}(\text{NO}_3)_2 \rightarrow 2 \text{HgO} + 4 \text{NO}_2 + \text{O}_2$$

- a. How many grams of mercury (II) oxide will be produced if 27.0 g of mercury (II) nitrate react?

$$\frac{27.0 \text{ g Hg}(\text{NO}_3)_2}{324.61 \text{ g Hg}(\text{NO}_3)_2} \times \frac{1 \text{ mol Hg}(\text{NO}_3)_2}{2 \text{ mol Hg}(\text{NO}_3)_2} \times \frac{2 \text{ mol HgO}}{2 \text{ mol Hg}(\text{NO}_3)_2} \times \frac{216.59 \text{ g HgO}}{1 \text{ mol HgO}} = 18.0 \text{ g HgO}$$

- b. How many moles of oxygen gas will be produced if 3.5 g of nitrogen dioxide are produced?

$$\frac{3.5 \text{ g NO}_2}{46.01 \text{ g NO}_2} \times \frac{1 \text{ mol NO}_2}{4 \text{ mol NO}_2} \times \frac{1 \text{ mol O}_2}{4 \text{ mol NO}_2} = 0.019 \text{ mol O}_2$$

- c. When 12.0 g of mercury (II) nitrate are decomposed in the lab, it is found that 7.56 g of mercury (II) oxide are produced. What is the actual yield, the theoretical yield, and the percent yield?

$$\frac{12.0 \text{ g Hg}(\text{NO}_3)_2}{524.61 \text{ g Hg}(\text{NO}_3)_2} \times \frac{1 \text{ mol Hg}(\text{NO}_3)_2}{2 \text{ mol Hg}(\text{NO}_3)_2} \times \frac{2 \text{ mol HgO}}{2 \text{ mol Hg}(\text{NO}_3)_2} \times \frac{216.59 \text{ g HgO}}{1 \text{ mol HgO}} = 270.6 \text{ g HgO}$$

$$\frac{7.56 \text{ g}}{270.6 \text{ g}} \times 100 = 94.49\%$$

4. If 5.6 g of copper (II) oxide are reacted with 8.6 g of hydrogen according to the following BALANCED reaction, how many grams of copper metal will be produced?
- $$\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$$

$$\frac{5.6 \text{ g CuO}}{79.56 \text{ g CuO}} \times \frac{1 \text{ mol CuO}}{1 \text{ mol CuO}} \times \frac{1 \text{ mol Cu}}{1 \text{ mol CuO}} \times \frac{63.56 \text{ g Cu}}{1 \text{ mol Cu}} = 4.5 \text{ g Cu}$$

$$\frac{8.6 \text{ g H}_2}{2.02 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1 \text{ mol H}_2} \times \frac{1 \text{ mol Cu}}{1 \text{ mol H}_2} \times \frac{63.56 \text{ g Cu}}{1 \text{ mol Cu}} = 270.6 \text{ g Cu}$$

The limiting reactant is CuO and the excess reactant is H₂

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