

60 What is the atomic radius of Carbon?

70pm

90 On each table, which element has the smallest atomic radius? What about the largest atomic radius?

Smallest H, N

Largest Rb, Ne

108 For the Group 1 table, what is the general trend of atomic radius vs. position?

Atomic radius increases as you go down group

110 For the Period 2 table, what is the general trend of atomic radius vs. position?

Atomic radius decreases as you move across period

120 Are there any exceptions that you noticed for either of the trends?

Noble gases are significantly larger

130 Speculate with your group why this trend-exception occurs. We will discuss as a class and why this exception occurs.

* Can discuss Van der Waals radius vs. Atomic radius & why Van der Waals radius is used for noble gases

Information: Generally, atomic radius increases as you go down and to the left on the periodic table.

As you go down on the periodic table, the energy level of the outer-most electron increases, which requires the electrons to be further from the nucleus.

As you go further to the right on the periodic table, the energy level of the outer-most electron doesn't change significantly. The effective nuclear charge does steadily increase though, which allows the nucleus to have a stronger force on the outer-most electrons of the atom.

Pogil Ap Chemistry Activities Answers

Shawn R. Simonson



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