

## PHYSICAL: Kinetic Molecular Theory

### Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are not attracted to each other).
- At standard atmospheric conditions, all gases have the same average kinetic energy.
- At standard atmospheric conditions, all gases have the same average molecular mass.

### Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three different phases of matter.



Figure 1: Representation of particles in different phases of matter.

At standard atmospheric conditions, all gases have the same average kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are not attracted to each other).

### Key Concepts

- At standard atmospheric conditions, all gases have the same average kinetic energy, and are made up of molecules that behave as independent particles (the particles of the gases are not attracted to each other).
- At standard atmospheric conditions, all gases have the same average molecular mass.
- At standard atmospheric conditions, all gases have the same average molecular mass.

Table 1: Representation of particles in different phases of matter.

	GAS	LIQUID	CRYSTALLINE SOLID
How do the particles move?	Particles move in all directions, randomly, and are not attracted to each other.	Particles move in all directions, randomly, and are not attracted to each other.	Particles move in all directions, randomly, and are not attracted to each other.
How do the particles move?	Particles move in all directions, randomly, and are not attracted to each other.	Particles move in all directions, randomly, and are not attracted to each other.	Particles move in all directions, randomly, and are not attracted to each other.

# Pogil Answer Kinetic Molecular Theory

**M Mosston**



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