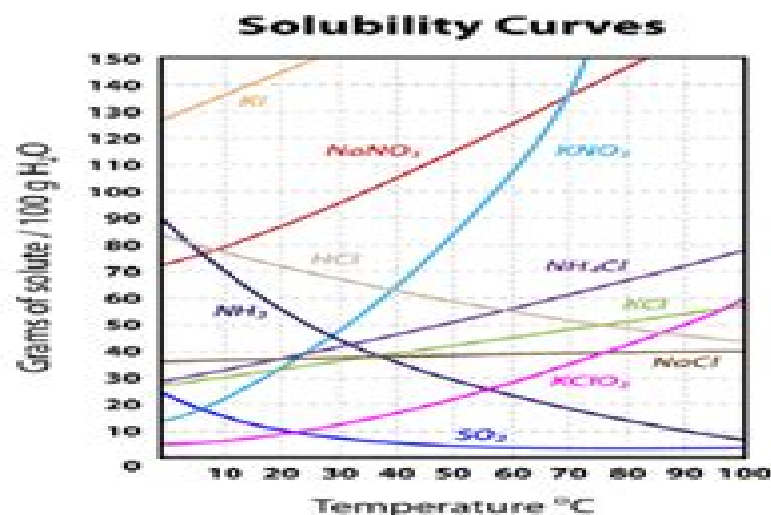


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 ml.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO_3 at 10°C , 40°C , & 80°C ?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH_3 can you dissolve in water at 10°C , 30°C , & 90°C ?
@ 10°C = ~90 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C , 50°C , & 90°C ?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C ?
About 40 grams
- Except for KCl What substance is the most soluble in water at 20°C ?
 NaNO_3
- Except for KCl what substance is the most soluble in water at 72°C ?
 KNO_3
- Look up The boiling points for SO_2 , HCl , and NH_3 and write them below:
B.P. SO_2 = -10°C
B.P. HCl = -85°C
B.P. NH_3 = -33°C
- Based on the boiling points you looked up, are SO_2 , HCl , & NH_3 solids, liquids, or gases at 0°C ? (and anywhere else on the solubility curve)
gases

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Table of Contents Unit 8 Solubility And Solubility Curve Answers

1. Understanding the eBook Unit 8 Solubility And Solubility Curve Answers
 - The Rise of Digital Reading Unit 8 Solubility And Solubility Curve Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 8 Solubility And Solubility Curve Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 8 Solubility And Solubility Curve Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 8 Solubility And Solubility Curve Answers
 - Personalized Recommendations

- Unit 8 Solubility And Solubility Curve Answers User Reviews and Ratings
- Unit 8 Solubility And Solubility Curve Answers and Bestseller Lists
- 5. Accessing Unit 8 Solubility And Solubility Curve Answers Free and Paid eBooks
 - Unit 8 Solubility And Solubility Curve Answers Public Domain eBooks
 - Unit 8 Solubility And Solubility Curve Answers eBook Subscription Services
 - Unit 8 Solubility And Solubility Curve Answers Budget-Friendly Options
- 6. Navigating Unit 8 Solubility And Solubility Curve Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 8 Solubility And Solubility Curve Answers Compatibility with Devices
 - Unit 8 Solubility And Solubility Curve Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 8 Solubility And Solubility Curve Answers
 - Highlighting and Note-Taking Unit 8 Solubility And Solubility Curve Answers
 - Interactive Elements Unit 8 Solubility And Solubility Curve Answers
- 8. Staying Engaged with Unit 8 Solubility And Solubility Curve Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 8 Solubility And Solubility Curve Answers
- 9. Balancing eBooks and Physical Books Unit 8 Solubility And Solubility Curve Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 8 Solubility And Solubility Curve Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 8 Solubility And Solubility Curve Answers
 - Setting Reading Goals Unit 8 Solubility And Solubility Curve Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 8 Solubility And Solubility Curve Answers
 - Fact-Checking eBook Content of Unit 8 Solubility And Solubility Curve Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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