## Acids, Bases and Salts

Reaction of acids with oxides of metals: When a metal oxide reacts with an acid it forms corresponding salt and water. Metallic oxides are basic oxides.

Metal oxide + Dilute acid ---> Salt + Water

$CaO_{(8)} + 2HCl_{(aq)}$	$> CaCl_{2(aq)} + H_2O_{(0)}$	
Calcium	Calcium	Water
oxide	chloride	

$$ZnO_{(s)} + 2HCl_{(aq)} \longrightarrow ZnCl_2 + H_2O_{(l)}$$
  
Zinc oxide Zinc chloride

#### Reactivity of base with non-metallic oxides:

When a non-metallic oxide reacts with a base, it forms salt and water. Nonmetallic oxides are acidic in nature.

Non-metallic oxide + Base --- Salt + Water

$$CO_{3(g)} + 2NaOH_{(aq)} \longrightarrow Na_2CO_{3(aq)} + H_2O_{(l)}$$
  
Carbon Sodium  
dioxide carbonate

$$CO_{2(g)} + 2KOH_{(aq)} \longrightarrow K_2CO_{3(aq)} + H_2O_{(b)}$$
  
Potassium carbonate

$$SO_{3(g)} + 2NaOH_{(aq)} \longrightarrow Na_2SO_{4(aq)} + H_2O_{(l)}$$
  
Sulphur trioxide Sodium sulphate

$$ZnO_{(s)} + 2NaOH_{(sq)} \longrightarrow Na_2ZnO_2 + H_2O_{(l)}$$
  
Sodium zincate

#### Salts

A salt is a combination of an anion of an acid and a cation of a base. Examples – KCl, NaNO<sub>3</sub>, CaSO<sub>4</sub>, etc. Salts are the ionic compounds which are produced after the neutralization reaction between acid and base. Salts are electrically neutral.

# **Unit Acids Bases And Salts Answers**

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