



Unit: Stoichiometry **Answer key** "Mole-Mole Calculations" - WS #1

Directions: For each of the following problems, show your balanced chemical equation and all of your calculations.

1. How many moles of potassium oxide (K_2O) will be formed when 1.52 moles of potassium reacts with oxygen according to the following reaction:



$$1.52 \text{ mol } K \times \frac{2 \text{ mol } K_2O}{4 \text{ mol } K} = 0.760 \text{ mol } K_2O$$

How many moles of copper are needed to react with sulfur to produce 0.25 moles of copper(I) sulfide according to the following reaction:



$$0.25 \text{ mol } Cu_2S \times \frac{2 \text{ mol } Cu}{1 \text{ mol } Cu_2S} = 0.50 \text{ mol } Cu$$

2. How many moles of sodium are needed to react with sulfuric acid to produce 3.75 moles of sodium sulfate according to the following (unbalanced) reaction:



$$3.75 \text{ mol } Na_2SO_4 \times \frac{2 \text{ mol } Na}{1 \text{ mol } Na_2SO_4} = 7.50 \text{ mol } Na$$

3. How many moles of calcium oxide will be produced when 1.6 moles of Iron(III)oxide reacts with calcium phosphate according to the following (unbalanced) reaction:



$$1.6 \text{ mol } Fe_2O_3 \times \frac{3 \text{ mol } CaO}{1 \text{ mol } Fe_2O_3} = 4.8 \text{ mol } CaO$$

4. How many moles of oxygen are produced when 5.12 moles of magnesium chlorate decomposes into magnesium chloride and oxygen?

Word equation: **magnesium chlorate \rightarrow magnesium chloride + oxygen**

Formula equation: **$Mg(ClO_3)_2 \rightarrow MgCl_2 + 3O_2$**
(must be balanced!)

$$5.12 \text{ mol } Mg(ClO_3)_2 \times \frac{3 \text{ mol } O_2}{1 \text{ mol } Mg(ClO_3)_2} = 15.4 \text{ mol } O_2$$

5. How many moles of Chromium(III) nitrate are produced when Chromium reacts with 0.85 moles of Lead(IV) nitrate to produce chromium(III) nitrate and lead?

Word equation: **chromium + lead (IV) nitrate \rightarrow chromium (III) nitrate + lead**

Formula equation: **$4Cr + 3Pb(NO_3)_4 \rightarrow 4Cr(NO_3)_3 + 3Pb$**
(must be balanced!)

$$0.85 \text{ mol } Pb(NO_3)_4 \times \frac{4 \text{ mol } Cr(NO_3)_3}{3 \text{ mol } Pb(NO_3)_4} = 1.1 \text{ mol } Cr(NO_3)_3$$

Unit Stoichiometry Problems Answers

JS Bruner



Unit Stoichiometry Problems Answers:

This book delves into Unit Stoichiometry Problems Answers. Unit Stoichiometry Problems Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Unit Stoichiometry Problems Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Unit Stoichiometry Problems Answers
 - Chapter 2: Essential Elements of Unit Stoichiometry Problems Answers
 - Chapter 3: Unit Stoichiometry Problems Answers in Everyday Life
 - Chapter 4: Unit Stoichiometry Problems Answers in Specific Contexts
 - Chapter 5: Conclusion
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3. In chapter 2, this book will delve into the foundational concepts of Unit Stoichiometry Problems Answers. This chapter will elucidate the essential principles that need to be understood to grasp Unit Stoichiometry Problems Answers in its entirety.
4. In chapter 3, this book will examine the practical applications of Unit Stoichiometry Problems Answers in daily life. The third chapter will showcase real-world examples of how Unit Stoichiometry Problems Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Unit Stoichiometry Problems Answers in specific contexts. This chapter will explore how Unit Stoichiometry Problems Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Unit Stoichiometry Problems Answers. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Unit Stoichiometry Problems Answers.

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web il est là c est une bougie le bonheur ris pas trop fort d ailleurs tu risques de l éteindre on l veut le bonheur on l veut ouais tout le monde veut l atteindre mais il fait pas de bruit le bonheur non il fait pas de bruit non il n en fait pas c est con le bonheur ouais car c est souvent après qu on sait qu il était là