

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{2.8\text{g}}{4.0\text{g}} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.}$$

$$1.55\text{mol CaCO}_3 \times \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} = 1.55\text{mol CaCl}_2$$

$$250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{E.R.}$$

$$6.857\text{mol HCl} \times \frac{1\text{mol CaCl}_2}{2\text{mol HCl}} = 3.43\text{mol CaCl}_2$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{142.0\text{g}}{172.02\text{g}} \right] \times 100 = 82.55\%$$

Unit 12 Stoichiometry Answers

Nancy Konigsberg Kerner

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

Unit 12 Stoichiometry Answers:

Chemistry Class - XII Model Paper Chapter wise Question Answer With Marking Scheme 2023- SBPD

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Conceptual Chemistry Volume-I For Class XII S.K. Jain & Shailesh K. Jain, Conceptual Chemistry Volume I For Class XII

Caveman Chemistry Kevin M. Dunn, 2003 Half a million years ago our ancestors learned to make fire from scratch They crafted intricate tools from stone and brewed mind altering elixirs from honey Their descendants transformed clay into pottery wool into clothing and ashes into cleansers In ceramic crucibles they won metal from rock the metals lead to colored glazes and glass Buildings of brick and mortar enshrined books of parchment and paper Kings and queens demanded ever more colorful clothing and accessories in order to out class clod hoppers and call girls Kingdoms rose and fell by the power of saltpeter sulfur and charcoal And the demands of everyday folk for glass and paper and soap stimulated the first round of chemical industrialization From sulfuric acid to sodium carbonate From aniline dyes to analgesic drugs From blasting powder to fertilizers and plastics In a phrase From Caveman to Chemist Your guides on this journey are the four alchemical elements Fire Earth Air and Water These archetypical characters deliver first hand accounts of the births of their respective technologies The spirit of Fire for example was born in the first creature to cultivate the flame This spirit passed from one person to another from one generation to another from one millennium to another arriving at last in the pages of this book The spirit of Earth taught folks to make tools of stone the spirit of Air imparted knowledge of units and the spirit of Water began with the invention of spirits Having traveled the world from age to age who can say where they will find their next home Perhaps they will find one in you

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