

## Balancing Equations Worksheet #2

Note – tough ones will have a coefficient above 3, all other problems use coefficients of 3 or fewer.

- 1) \_\_\_\_  $\text{Li}_3\text{PO}_4$  + \_\_\_\_  $\text{NaOH}$   $\rightarrow$  \_\_\_\_  $\text{LiOH}$  + \_\_\_\_  $\text{Na}_3\text{PO}_4$
- 2) \_\_\_\_  $\text{MgF}_2$  + \_\_\_\_  $\text{Li}_2\text{CO}_3$   $\rightarrow$  \_\_\_\_  $\text{MgCO}_3$  + \_\_\_\_  $\text{LiF}$
- 3) \_\_\_\_  $\text{P}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{P}_2\text{O}_3$
- 4) \_\_\_\_  $\text{RbNO}_3$  + \_\_\_\_  $\text{MgF}_2$   $\rightarrow$  \_\_\_\_  $\text{Mg}(\text{NO}_3)_2$  + \_\_\_\_  $\text{RbF}$
- 5) \_\_\_\_  $\text{AgNO}_3$  + \_\_\_\_  $\text{Cu}$   $\rightarrow$  \_\_\_\_  $\text{Cu}(\text{NO}_3)_2$  + \_\_\_\_  $\text{Ag}$
- 6) \_\_\_\_  $\text{CF}_4$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{CBr}_4$  + \_\_\_\_  $\text{F}_2$
- 7) \_\_\_\_  $\text{HCN}$  + \_\_\_\_  $\text{CuSO}_4$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{SO}_4$  + \_\_\_\_  $\text{Cu}(\text{CN})_2$
- 8) \_\_\_\_  $\text{GaF}_3$  + \_\_\_\_  $\text{Cs}$   $\rightarrow$  \_\_\_\_  $\text{CsF}$  + \_\_\_\_  $\text{Ga}$
- 9) \_\_\_\_  $\text{SrS}$  + \_\_\_\_  $\text{PtF}_2$   $\rightarrow$  \_\_\_\_  $\text{SrF}_2$  + \_\_\_\_  $\text{PtS}$
- 10) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_\_  $\text{NH}_3$
- 11) \_\_\_\_  $\text{LiF}$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{LiBr}$  + \_\_\_\_  $\text{F}_2$
- 12) \_\_\_\_  $\text{Pb}(\text{OH})_2$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_  $\text{PbCl}_2$
- 13) \_\_\_\_  $\text{GaBr}_3$  + \_\_\_\_  $\text{Na}_2\text{CO}_3$   $\rightarrow$  \_\_\_\_  $\text{NaBr}$  + \_\_\_\_  $\text{Ga}_2(\text{CO}_3)_3$  (a tough one)
- 14) \_\_\_\_  $\text{CH}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$
- 15) \_\_\_\_  $\text{Li}_3\text{PO}_4$  + \_\_\_\_  $\text{CaCl}_2$   $\rightarrow$  \_\_\_\_  $\text{LiCl}$  + \_\_\_\_  $\text{Ca}_3(\text{PO}_4)_2$  (also tough)
- 16) \_\_\_\_  $\text{Na}$  + \_\_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_\_  $\text{NaCl}$
- 17) \_\_\_\_  $\text{Ga}$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2$  + \_\_\_\_  $\text{GaCl}_3$  (last tough one)
- 18) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{F}_2$   $\rightarrow$  \_\_\_\_  $\text{NF}_3$
- 19) \_\_\_\_  $\text{SO}_2$  + \_\_\_\_  $\text{Li}_2\text{Se}$   $\rightarrow$  \_\_\_\_  $\text{SSe}_2$  + \_\_\_\_  $\text{Li}_2\text{O}$
- 20) \_\_\_\_  $\text{NH}_3$  + \_\_\_\_  $\text{H}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_  $(\text{NH}_4)_2\text{SO}_4$

# Unit 7 Balancing Equations 2 Answers

**John Hindmarsh, Alasdair Renfrew**



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