

Q1.

This question is about reversible reactions and chemical equilibrium.

(a) Reversible reactions can reach equilibrium in a closed system.

(i) What is meant by a closed system?

(1)

(ii) Explain why, when a reversible reaction reaches equilibrium, the reaction appears to have stopped.

(2)

(b) In the Haber process, the reaction of nitrogen with hydrogen to produce ammonia is reversible.



(i) Name a natural resource from which hydrogen is produced.

(1)

(ii) The Haber process uses a catalyst to speed up the reaction.

Explain how a catalyst speeds up a reaction.

(2)

(iii) What happens to the amount of ammonia produced at equilibrium if the pressure is increased?

Give a reason for your answer.

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JE Gale



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Table of Contents Nscience Question Paper

1. Understanding the eBook Nscience Question Paper
 - The Rise of Digital Reading Nscience Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Nscience Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nscience Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nscience Question Paper
 - Personalized Recommendations

- Nscience Question Paper User Reviews and Ratings
- Nscience Question Paper and Bestseller Lists
- 5. Accessing Nscience Question Paper Free and Paid eBooks
 - Nscience Question Paper Public Domain eBooks
 - Nscience Question Paper eBook Subscription Services
 - Nscience Question Paper Budget-Friendly Options
- 6. Navigating Nscience Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Nscience Question Paper Compatibility with Devices
 - Nscience Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nscience Question Paper
 - Highlighting and Note-Taking Nscience Question Paper
 - Interactive Elements Nscience Question Paper
- 8. Staying Engaged with Nscience Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nscience Question Paper
- 9. Balancing eBooks and Physical Books Nscience Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nscience Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nscience Question Paper
 - Setting Reading Goals Nscience Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nscience Question Paper
 - Fact-Checking eBook Content of Nscience Question Paper

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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