

Unit 2: Atoms and Elements  
AP Chemistry  
The Nuclear Atom

ALL of the answers to this worksheet can be **logically** figured out by looking at the **Schematic Diagrams for Various Atoms**, the **Periodic Table**, and **discussing** with your partners. All of the information you need is here somewhere. Determine each answer and be able to give convincing reasons for each answer. Good luck.

1. How many **protons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **6**
2. How many **neutrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **7**       $^{14}\text{C}$ ? **7**
3. How many **electrons** are found in  $^{12}\text{C}$ ? **6**       $^{13}\text{C}$ ? **6**       $^{14}\text{C}$ ? **7**
4. Based on the model,
  - a) what do all carbon atoms (and ions) have in common?  
**Same number and types of protons.**
  - b) what do all hydrogen atoms (and ions) have in common?  
**Same number and types of protons.**
5. What is the significance of the atomic number,  $Z$ , above each atomic symbol in the periodic chart?  
**It tells us the number of protons located in the nucleus of an atom of that particular element.**
6. What do all nickel ( $\text{Ni}$ ) atoms have in common?  
**Same number and types of protons.**
7. How is the mass number,  $A$ , (left-hand superscript next to the atomic symbol) determined?  
**By adding the protons and neutrons of an atom together.**
8. What structural feature is different in isotopes of a particular element?  
**The number of neutrons that are present.**
9.
  - a) What feature distinguishes a neutral atom from an ion?  
**The number of electrons that are present.**
  - b) How is the charge on an ion determined?  
**By subtracting the number of electrons from the number of protons in the atom.**
10. Where is most of the mass of an atom, within the nucleus or outside of the nucleus? Explain your reasoning.  
**Within the nucleus since that is where the protons and neutrons are located.**
11. Complete the following table:

Isotope	Atomic Number $Z$	Mass Number $A$	Number of electrons
$^{31}\text{P}$	15	31	15
$^{18}\text{O}$	8	18	8
$^{39}\text{Y}^{+}$	19	39	18
$^{58}\text{Ni}^{2+}$	28	58	26

# Nuclear Chemistry Answer Key

**MTG Learning Media**

A red, semi-circular graphic element with a gradient, positioned to the right of the MTG Learning Media text.

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## **Table of Contents Nuclear Chemistry Answer Key**

1. Understanding the eBook Nuclear Chemistry Answer Key
  - The Rise of Digital Reading Nuclear Chemistry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Nuclear Chemistry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nuclear Chemistry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nuclear Chemistry Answer Key
  - Personalized Recommendations
  - Nuclear Chemistry Answer Key User Reviews and Ratings
  - Nuclear Chemistry Answer Key and Bestseller Lists
5. Accessing Nuclear Chemistry Answer Key Free and Paid eBooks
  - Nuclear Chemistry Answer Key Public Domain eBooks
  - Nuclear Chemistry Answer Key eBook Subscription Services
  - Nuclear Chemistry Answer Key Budget-Friendly Options

6. Navigating Nuclear Chemistry Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Nuclear Chemistry Answer Key Compatibility with Devices
  - Nuclear Chemistry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Nuclear Chemistry Answer Key
  - Highlighting and Note-Taking Nuclear Chemistry Answer Key
  - Interactive Elements Nuclear Chemistry Answer Key
8. Staying Engaged with Nuclear Chemistry Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Nuclear Chemistry Answer Key
9. Balancing eBooks and Physical Books Nuclear Chemistry Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Nuclear Chemistry Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Nuclear Chemistry Answer Key
  - Setting Reading Goals Nuclear Chemistry Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Nuclear Chemistry Answer Key
  - Fact-Checking eBook Content of Nuclear Chemistry Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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