

Reactions of Amines and Epoxides

Preparation of Amines

Amines are colorless, basic, and have a strong, fishy odor. They are prepared by the reduction of nitriles.



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Epoxide Reactions and Mechanisms

Bond Types, Strength, Length, Polarity and Dipole Moment

Free-radical halogenation



There are two types of bonding: ionic bonding and covalent bonding.

• **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity (electron). Examples:



• **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of Atoms



Bond and Bonding electron



Bonding electron



There are two types of electronegativity

Electronegativity of elements are



Sigma and pi bonds

Covalent bonds are classified into sigma and pi bonds.

• **Sigma bonds** are formed by head-on (end-to-end) overlap of orbitals.

• **pi bonds** are formed by side-on overlap of orbitals.

• **One single bond is a sigma bond.**

• **Two multiple bonds consist of one sigma and two pi bonds.**



• **pi bonds are stronger than sigma bonds.**

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds are

classified into polar and nonpolar.



• If the difference in electronegativity is less than 0.5, the electrons are shared equally (nonpolar covalent bond). Examples are the C-C and C-H bonds.

• If the difference in electronegativity is between 0.5 and 1.7, the electrons are shared unequally (polar covalent bond). Examples are the C-O and C-N bonds.

• A difference of 1.7 or higher is ionic, where the electrons are not shared (ionic bond).

Dipole Moment

The polarity of a bond is described by the dipole moment (μ).

It is a vector quantity.



The direction of the dipole moment is from the positive end to the negative end.

• The dipole moment is a vector pointing from the less electronegative to the more electronegative atom.

• The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).

• **δ^- (delta minus)** indicates an electron-deficient atom.

• **δ^+ (delta plus)** indicates an electron-rich atom.

Molecular Dipole

To determine the molecular polarity, the vector sum of the dipole moments is calculated.

The geometry of the molecule is important. Symmetrical molecules are nonpolar, while asymmetrical molecules are polar.



• **Nonpolar molecules** have no net dipole moment.

Bond Length and Bond Strength

The bond length is the distance between the nuclei of the two atoms.

Bond	H-H	H-Cl	H-Br	H-I
Length	109 pm	127 pm	141 pm	154 pm
Strength	436 kJ/mol	431 kJ/mol	366 kJ/mol	298 kJ/mol

• **Stronger bonds** have shorter bond lengths and higher bond strengths.

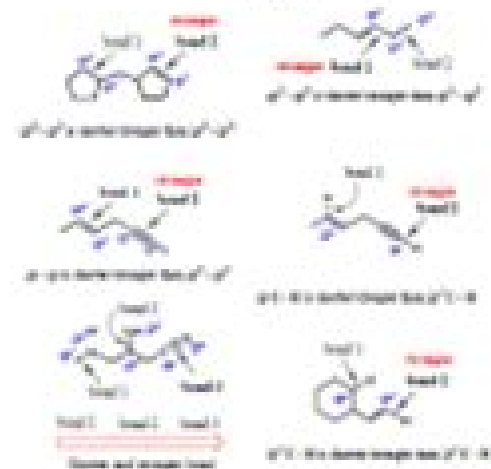


The bond length is the distance between the nuclei of the two atoms.

Average values of C-C and C-H bonds depending on hybridization			
C-C Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	154 pm	134 pm	121 pm
Strength	347 kJ/mol	461 kJ/mol	591 kJ/mol
Hybridization	sp^3-sp^3	sp^2-sp^2	$sp-sp$
C-H Bond	sp^3-sp^3	sp^2-sp^2	$sp-sp$
Length	109 pm	108 pm	106 pm
Strength	413 kJ/mol	461 kJ/mol	552 kJ/mol

Higher s character means stronger bond

Examples



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