

Chapter 12 STUDY GUIDE

Mixtures and Solutions

Section 12.1: Homogeneous and Heterogeneous Mixtures

Do your best to answer each question completely and correctly.

The van der Waals equation states that there are three factors that affect the behavior of a gas.

1. The volume of the gas molecules is not negligible compared to the volume of the container.
2. The attractive forces between gas molecules are not negligible.
3. The pressure of the gas is not negligible.
4. The volume of the gas molecules is not negligible compared to the volume of the container.
5. The pressure of the gas is not negligible.
6. The volume of the gas molecules is not negligible compared to the volume of the container.
7. The pressure of the gas is not negligible.
8. The volume of the gas molecules is not negligible compared to the volume of the container.
9. The pressure of the gas is not negligible.
10. The volume of the gas molecules is not negligible compared to the volume of the container.
11. The pressure of the gas is not negligible.
12. The volume of the gas molecules is not negligible compared to the volume of the container.
13. The pressure of the gas is not negligible.
14. The volume of the gas molecules is not negligible compared to the volume of the container.
15. The pressure of the gas is not negligible.
16. The volume of the gas molecules is not negligible compared to the volume of the container.
17. The pressure of the gas is not negligible.
18. The volume of the gas molecules is not negligible compared to the volume of the container.
19. The pressure of the gas is not negligible.
20. The volume of the gas molecules is not negligible compared to the volume of the container.

The van der Waals equation states that there are three factors that affect the behavior of a gas. The van der Waals equation states that there are three factors that affect the behavior of a gas.

Factor	Volume	Pressure	Temperature
1. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
2. The attractive forces between gas molecules are not negligible.	Volume	Pressure	Temperature
3. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
4. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
5. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
6. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
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8. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
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13. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
14. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
15. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
16. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
17. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
18. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
19. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
20. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature

Study Guide For Mixtures And Solutions



**Lawrie Ryan, Denise
Hernandez, Bermadee McKenzie-
Briscoe, Marsha Russell, Victor Joseph**

Study Guide For Mixtures And Solutions:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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