

Unit 2 Exam Study Guide

Use this guide to direct your attention toward the information to be covered by the test.

Biochemistry (Chapter 2)

1. Which type of energy moves matter?
Kinetic Energy
2. Which elements make up around 96% of the body's mass?
Carbon, Oxygen, Hydrogen, Nitrogen
3. What is the most abundant element in the human body?
Carbon
4. Which element is the simplest atom?
Hydrogen
5. Where are the three subatomic particles that make up the atom located?
Proton and Neutrons: Nucleus of Atom
Electrons: outside of Nucleus
6. What are the electrical charges of the subatomic particles?
Protons: +
Neutrons: No charge
Electrons: -
7. What happens when sodium reacts with chlorine to form salt?
Ionic Bond
8. Which type of bond is formed when electrons are shared equally?
Covalent Bond
9. How does a compound differ from a molecule?
Molecule: two or more of the same atoms binds together
Compound: two or more different atoms bind together
10. What is an example of a synthesis reaction?
 $A + B \rightarrow AB$
11. Why is water useful in body processes?
High Heat Capacity
Polarity/Solvent Properties
Chemical reactivity
Cushioning
12. Which numbers on the pH scale indicate an acid?
0-6.9
13. What is the definition of an inorganic molecule?
Molecules that do not have repeating carbon chains
14. Which type of compound is starch?
Carbohydrate (organic)
15. Which organic compounds serve as enzymes in catalyzing reactions?
Proteins
16. How do the nitrogenous bases of DNA pair up when in the double helix?
Adenine-Thymine
Guanine-Cytosine
17. How would you describe the function of enzymes?
Acts as a catalyst by lowering activation energy
18. Which element is present in proteins and nucleic acids but not in carbohydrates or lipids?
Nitrogen
19. How does the structure of RNA differ from that of DNA?
DNA is double stranded, RNA is single Stranded
RNA contains Uracil, DNA contains thymine
20. Be able to identify the four major classes of organic compounds by shape.
Check notes for structures
21. What is the outermost shell of electrons?
Valence Electrons
22. Which word describes molecules that make up living matter but do not contain carbon?
Inorganic
23. What are some examples of monosaccharides?
Glucose, galactose
24. What is the energy compound that supplies usable energy to all our cells?
ATP
25. Be familiar with the properties of water that are important to living things.
High Heat capacity: Water can absorb a lot of heat
Universal Solvent
Polar: uneven distribution of charges

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Sebastian Brünink

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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