

Chemistry 101 Final Exam Study Guide- Spring 2017

Acids and Bases

Acid: substance that provides H^+ ions in water

Base: substance that provides OH^- ions in water

Covalent Bonds

- Review: ionic bonds are formed by the positive charge of the cation being attracted to the negative charge of the anion
- Covalent bonds: bonds formed by sharing electrons between the atoms (not ions)
 - 1 electron from each atom (2 atoms per bond)
 - Electrons available for bonding are the un-paired electrons in orbitals (orbitals with only 1 electron in them)
- Molecule
 - A group of atoms held together by covalent bonds
- Atoms will bond with enough other atoms to complete the octet
 - But they can only form as many bonds as unpaired valence electrons
 - Hydrogen has 1 valence electron, so it can only form 1 bond, and will not achieve an octet
 - Boron has 3 valence electrons, so can only form 3 bonds, and will not achieve the octet (only 6 electrons)
 - There are always exceptions to the rules
 - Sulfur and phosphorus can extend the octet

Multiple Covalent Bonds between 2 atoms

- if the octet is not met, additional bonds may be formed
 - CO_2 (carbon dioxide)
 - C has 4 valence electrons, so 4 bonds possible
 - O has 6 valence electrons, 2 are unpaired, so 2 bonds are possible

PRACTICE

IONS:

- Sodium
 - Symbol: Na
 - Valence: 1
 - Electron dot: Na^{\bullet}
 - Ionic dot: Na^+ (loses its electron to get a 1+ charge)
- Magnesium
 - Symbol: Mg
 - 2 valence e^-
 - Loses 2 electrons to get 2+ charge
- Oxygen
 - Symbol: O
 - Valence: 6
 - Will gain 2 electrons for a 2- charge

Cannot have a covalent bond with ions, just atoms

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Liying Dong



Spring Final Exam Study Guide Chemistry:

Unveiling the Magic of Words: A Review of "**Spring Final Exam Study Guide Chemistry**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Spring Final Exam Study Guide Chemistry**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect on the souls of its readers.

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