

# Naming Molecular Compounds

How are the chemical formula and name of a molecular compound related?

## Why?

When you began chemistry class this year, you probably already knew that the chemical formula for carbon dioxide was  $\text{CO}_2$ . Today you will find out why  $\text{CO}_2$  is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide ( $\text{CO}$ , a poisonous, deadly gas) and carbon dioxide ( $\text{CO}_2$ , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

## Model 1 – Molecular Compounds

Molecular Formula	Number of Atoms of First Element	Number of Atoms of Second Element	Name of Compound
$\text{ClF}$	1	1	Chlorine monofluoride
$\text{ClF}_5$	1	5	Chlorine pentafluoride
$\text{CO}$	1	1	Carbon monoxide
$\text{CO}_2$	1	2	Carbon dioxide
$\text{Cl}_2\text{O}$	2	1	Dichlorine monoxide
$\text{PCl}_5$	1	5	Phosphorus pentachloride
$\text{N}_2\text{O}_5$	2	5	Dinitrogen pentoxide

1. Fill in the table to indicate the number of atoms of each type in the molecular formula.

2. Examine the molecular formulas given in Model 1 for various molecular compounds.

a. How many different *elements* are present in each compound shown?

*Each compound has two elements.*

b. Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?

*The compounds have nonmetals with nonmetals.*

c. Based on your answer to b, what type of bonding must be involved in molecular compounds?

*A covalent bond must be involved.*

3. Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.

*$\text{ClF}$  and  $\text{ClF}_5$ : Chlorine fluoride does not specify that the two compounds have different ratios of chlorine and fluorine.*

4. Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

mono – *one* –  $\text{CO}$  (carbon monoxide) and  $\text{Cl}_2\text{O}$  (dichlorine monoxide)

di – *two* –  $\text{CO}_2$  (carbon dioxide) and  $\text{Cl}_2\text{O}$  (dichlorine monoxide)

penta – *five* –  $\text{ClF}_5$  (chlorine pentafluoride) and  $\text{PCl}_5$  (phosphorus pentachloride)

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