

Section 18.4 Neutralization*In your textbook, read about neutralization and titration.*

For each item in Column A, write the letter of the matching item in Column B.

Column A

- _____ 1. A chemical dye that changes color based on the pH of a solution
- _____ 2. A method for using a neutralization reaction to determine the concentration of a solution
- _____ 3. A reaction in which an acid and a base react to produce a salt and water
- _____ 4. A solution of known concentration
- _____ 5. An ionic product of an acid-base reaction
- _____ 6. The point in a titration in which an indicator changes color
- _____ 7. The stoichiometric point of a titration

Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

| Acid | Base | Salt formula | Salt name |
|-----------------------------|--------------------------|--------------|--------------------|
| 8. HCl | KOH | KCl | potassium chloride |
| 9. H_2SO_4 | $\text{Mg}(\text{OH})_2$ | | |
| 10. H_3PO_4 | NaOH | | |
| 11. HNO_3 | $\text{Fe}(\text{OH})_3$ | | |
| 12. H_3PO_4 | $\text{Ca}(\text{OH})_2$ | | |

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

Sequence of Steps

- _____ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- _____ 14. Fill a buret with the standard solution.
- _____ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

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