

# < KINETIC THEORY >

Introduction: - (Microscopic properties)

(John Dalton) → Atom → 1803

In 1873 → [Maxwell & Boltzmann]

Ideal Gas → [Kinetic theory of gases]

Atoms/molecules constantly motions  
They are not in rest.

< We study Average properties of gases >

Average relax time 'T'

- Average time period b/w two successive collisions.

$$T_m = T_1 + T_2 + \dots + T_n$$

(mean)  $n$

< In 1661 > → [Boyle's Law]:-

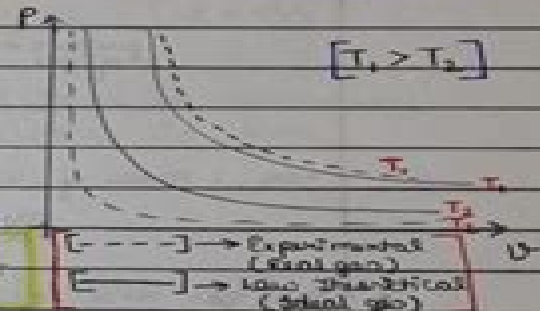
- Keeping temperature constant, pressure of a given mass of gas varies inversely with volume.

$$(P \propto 1/V)$$

$$\Rightarrow PV = \text{constant}$$

$$\text{or } [P_1 V_1 = P_2 V_2]$$

- At high temperature Real gas and Ideal gas apparently behave same. (High temp. pin Boyle's law real, ideal same gas. In low temp. real gas, otherwise diff. Ideal gas.)



< In 1783 > → [Charles's Law]:-

- For a fixed mass pressure, the vol. of a gas is directly proportional to its absolute temperature 'T'.

$$(V \propto T)$$

$$\Rightarrow \frac{V}{T} = \text{constant}$$

$$\text{or } \left[ \frac{V_1}{T_1} = \frac{V_2}{T_2} \right]$$

- At high temperature law holds for all gas (Real + Ideal).

Combining both Boyle's and Charles's law

$$(T = \text{constant}) V \propto \frac{1}{P} \quad , \quad V \propto T \quad (P = \text{constant})$$

$$\Rightarrow V \propto \frac{T}{P} \quad \text{or} \quad \frac{PV}{T} = \text{constant}$$

$$\text{or} \quad \frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

[Equation of State]

# Study Guide For Kinetic Theory

**Arshad Iqbal**



## Study Guide For Kinetic Theory:

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Measurement of Physical Quantities Questions Chapter 11 Measurement of Temperature Questions Chapter 12  
Measurements Questions Chapter 13 Melting and Boiling Questions Chapter 14 Pressure Questions Chapter 15 Properties  
and Mechanics of Matter Questions Chapter 16 Simple Kinetic Theory of Matter Questions Chapter 17 Sound Questions  
Chapter 18 Speed Velocity and Acceleration Questions Chapter 19 Temperature Questions Chapter 20 Thermal Energy  
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### **Table of Contents Study Guide For Kinetic Theory**

1. Understanding the eBook Study Guide For Kinetic Theory
  - The Rise of Digital Reading Study Guide For Kinetic Theory
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Kinetic Theory
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For Kinetic Theory
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Kinetic Theory
  - Personalized Recommendations
  - Study Guide For Kinetic Theory User Reviews and Ratings

- Study Guide For Kinetic Theory and Bestseller Lists
- 5. Accessing Study Guide For Kinetic Theory Free and Paid eBooks
  - Study Guide For Kinetic Theory Public Domain eBooks
  - Study Guide For Kinetic Theory eBook Subscription Services
  - Study Guide For Kinetic Theory Budget-Friendly Options
- 6. Navigating Study Guide For Kinetic Theory eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide For Kinetic Theory Compatibility with Devices
  - Study Guide For Kinetic Theory Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide For Kinetic Theory
  - Highlighting and Note-Taking Study Guide For Kinetic Theory
  - Interactive Elements Study Guide For Kinetic Theory
- 8. Staying Engaged with Study Guide For Kinetic Theory
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For Kinetic Theory
- 9. Balancing eBooks and Physical Books Study Guide For Kinetic Theory
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For Kinetic Theory
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide For Kinetic Theory
  - Setting Reading Goals Study Guide For Kinetic Theory
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide For Kinetic Theory
  - Fact-Checking eBook Content of Study Guide For Kinetic Theory
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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