

(Actual Test total points = 72)

1.12a<sup>1</sup> One atom of an element would most likely weigh closest to 2 grams:

- a)  $5 \times 10^{23}$       b)  $1 \times 10^{23}$       c)  $11 \times 10^2$       d)  $0.01 \times 10^{-23}$       e)  $0.000 \times 10^{-23}$

b<sup>1</sup> The number of atoms in 1.0 moles of water is closest to:

- a)  $3 \times 10^{23}$       b)  $1 \times 10^{23}$       c)  $11 \times 10^2$       d)  $0.01 \times 10^{-23}$       e)  $0.000 \times 10^{-23}$

2.3<sup>1</sup> Why do chemists need to use the mole unit?2.4<sup>2</sup> Identify the reaction type for each equation below:4.1<sup>3</sup> Be able to calculate molar mass. SHOW ALL WORK; LABEL WITH CORRECT UNITS!

a. Na

b. Fe

c.  $(\text{CH}_3)_2\text{CHOH}$ d.  $\text{F}_2$ e.  $\text{P}_4$ f.  $\text{Fe}(\text{OH})_3$ 4.4<sup>4</sup> Balance the following equations:

4.6<sup>5</sup> Consider the reaction:  $\text{L} + \text{M} \rightarrow \text{X} + \text{Z}$ . 57.0 g of L and 48.0 g of M are placed in a flask. After they are done reacting, L is all used up, 12.0 g of M remain, and 79.0 g of X were produced. What mass of Z must have been produced during the reaction? The data is summarized in the table below.

	L	M	X	Z
Mass before reaction	57.0 g	48.0 g	0 g	0 g
Mass after reaction	0 g	12.0 g	79.0 g	???

Answer: \_\_\_\_\_

Show your work:

# Skill Practice 33 Chemistry

**Joanne Capper**



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web mar 16 2015 ok so i figured out that the relay is in a box next to the radiator putting power to that relay engages the ac clutch which is just below it clutch works relay works but when i press the ac button the clutch and fans do not engage

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web nov 9 2021 here you will find the locations and photos of the fuse blocks separately we note the elements responsible for the cigarette lighter and fuel pump in the engine compartment

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