

NAME: _____

Equilibrium Practice

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1. The reaction below has a negative ΔH . Write the equilibrium constant expression for this reaction.
2. Consider the following chemical equation: $2\text{H}_2\text{O}(g) \rightleftharpoons 2\text{H}_2(g) + \text{O}_2(g)$. In equilibrium, the concentration of H_2 is 0.02 M, the concentration of H_2O is 0.01 M, and the concentration of O_2 is 0.01 M. Calculate the equilibrium constant, K .
3. At a certain pressure, 10 g of each of the gases, chlorine and hydrogen, were placed in a 10.0 L container containing the following equation: $\text{H}_2 + \text{Cl}_2 \rightleftharpoons 2\text{HCl}$. At equilibrium, the concentration of each gas was 0.01 M. Calculate the equilibrium constant, K .
4. A 10.0 L flask contains a gas mixture at a 1.00 M total of gas. The flask is heated to 100°C and the following chemical equation is shown: $\text{H}_2 + \text{Cl}_2 \rightleftharpoons 2\text{HCl}$.
5. Consider the following reaction: $\text{PCl}_5(g) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)$. If the initial concentration of PCl_5 is 1.00 M, what is the equilibrium concentration of Cl_2 if the concentration of PCl_3 is 0.50 M?

Skill Practice 45 Equilibrium Practice Chemistry Inquiry

Ying-Ying Zheng



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