

SINGLE REPLACEMENT REACTIONS WORKSHEET

- Using the activity series table answer the following questions:
 - list two metals that will react with Fe^{2+} ions _____
 - list two metals that will lose electrons to Pb^{2+} _____
 - list two ions that will take electrons from Zn _____
 - list two metals that will react with HCl _____
 - identify the best oxidizing agent (i.e. electron grabber) _____
- Using the activity series, how do you decide if a reaction will or will not take place between two species?
- Describe what you would see when a piece of Mg metal is placed in an aqueous solution of copper(II) sulfate. Write the net ionic equation for the reaction that takes place.
- What would happen to a silver dollar when placed in a 3 M HCl solution? Justify your answer.
- Given an ionic equation, how can you identify the spectator ions of a reaction?
- Do aluminum pop cans make a suitable containers for hydrochloric acid solutions? Write an equation that supports your answer.
- Rank order the following species in order of their **increasing** ability to capture electrons: Pb^{2+} , I_2 , Al^{3+} , Ba^{2+} , O_2

- Rank order the following species in order of their **increasing** ability to lose electrons: Ni, Cu^+ , Cl, K, Mg

- What happens when a piece of calcium metal is placed in water? Write the molecular equation that supports your prediction.

Skill Practice 28 Answers On Single Replacement

M Mosston



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