

Study Question 2d

New Chemical Reaction

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Chemical Thermodynamic Relationships

1. Standard reaction has worked the following reactions:

(a) $\text{Acetylene} + \text{hydrogen} \rightarrow \text{ethane}$ with standard enthalpy of formation: $\Delta H_f^\circ(\text{ethane}) = -84.68 \text{ kJ/mol}$



(b) $\text{Ethane} + \text{oxygen} \rightarrow \text{carbon dioxide} + \text{water}$ with standard enthalpy of formation: $\Delta H_f^\circ(\text{ethane}) = -84.68 \text{ kJ/mol}$



2. Standard free energy values:

(a) $\Delta G_f^\circ(\text{C}_2\text{H}_2, \text{g}) = 65.69 \text{ kJ/mol}$, $\Delta G_f^\circ(\text{H}_2, \text{g}) = 0$, $\Delta G_f^\circ(\text{C}_2\text{H}_6, \text{g}) = -32.82 \text{ kJ/mol}$

(b) $\Delta G_f^\circ(\text{H}_2\text{O}, \text{l}) = -237.13 \text{ kJ/mol}$, $\Delta G_f^\circ(\text{CO}_2, \text{g}) = -394.36 \text{ kJ/mol}$

Combustion Processes

3. Standard combustion reactions for the following compounds:

(a) C_2H_2 : $\text{C}_2\text{H}_2(\text{g}) + 5/2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ $\Delta H_c^\circ = -1300 \text{ kJ/mol}$

(b) C_2H_6 : $\text{C}_2\text{H}_6(\text{g})$

Equilibrium and Thermodynamic Relationships

4. Write an equation for the decomposition process:



5. Write an equation for the reaction of carbon monoxide with oxygen:



Given: $\Delta H_f^\circ(\text{CO}, \text{g}) = -110.5 \text{ kJ/mol}$, $\Delta H_f^\circ(\text{CO}_2, \text{g}) = -393.5 \text{ kJ/mol}$, $\Delta H_f^\circ(\text{H}_2\text{O}, \text{l}) = -285.8 \text{ kJ/mol}$

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