

REVIEW QUESTIONS FOR FINAL EXAM (2nd Semester)

SECTION 12.1

Heat Transfer: $Q = mc\Delta T$ or $Q = mc(T_f - T_i)$

Specific Heat of Common Substances			
Material	Specific Heat $J/kg \cdot ^\circ C$	Material	Specific Heat $J/kg \cdot ^\circ C$
aluminum	900	lead	130
brass	376	nickel	210
carbon	712	silver	235
copper	385	zinc	388
glass	840	water	4180
ice	2090	oil	2000
iron	450		

1) Two solid objects are in contact as shown below.



Heat will always flow from object A to object B if object B has a lower

- a) mass b) total energy c) specific heat
d) temperature e) none of the answers

For questions 2-3: The diagram below shows two solids with their temperatures ($A=50^\circ C$, $B=80^\circ C$).



2) What occurs when the two solids are placed in contact with each other as shown?

- a) Heat energy flows from A to B. Solid A cools down.
b) Heat energy flows from A to B. Solid A gets hotter.
c) Heat energy flows from B to A. Solid B cools down.
d) Heat energy flows from B to A. Solid B gets hotter.

3) The net flow of heat energy will stop from one object to the other when they both have the same

- a) energy **b) temperature**
c) specific heat d) all the answers are correct

4) How much heat is transferred when a 0.2 kg piece of copper's temperature is raised from $50^\circ C$ to $80^\circ C$?

$$Q = mc\Delta T$$

$$Q = (0.2 \text{ kg})(385 \text{ J/kg} \cdot ^\circ C)(80^\circ C - 50^\circ C) = \boxed{2310 \text{ J}}$$

5) A 0.5 kg block of lead cools down from $30^\circ C$ to $20^\circ C$. How much heat was transferred?

$$Q = (0.5 \text{ kg})(130 \text{ J/kg} \cdot ^\circ C)(20^\circ C - 30^\circ C)$$

$$\boxed{Q = -650 \text{ J}}$$

6) The heat transfer of a substance is 9400 J. The mass of this substance is 2 kg and its temperature rises from $80^\circ C$ to $100^\circ C$. Based on the specific heat that you will calculate for this substance, what is it?

$$C = \frac{Q}{m\Delta T} = \frac{9400 \text{ J}}{(2 \text{ kg})(100^\circ C - 80^\circ C)} = 235 \text{ J/kg} \cdot ^\circ C$$

SILVER

7) From the choices below, which material would be better for a pan handle?

- a) brass 376 b) silver 235 c) copper 385
d) iron 450 e) zinc 388 f) lead 130

8) Provide a brief description of absolute zero and their temperatures in Kelvin (K) and degrees Celsius ($^\circ C$).

The lowest temperature when all particle motion stops. $T_C = -273^\circ C$, $T_K = 0 \text{ K}$.

9) The boiling point of water is $100^\circ C$. The boiling point of water in Kelvin (K) will be

$$T_K = 273 + T_C = 273 + 100 = \boxed{373 \text{ K}}$$

10) A 1-liter of water taken out of a 5-gallon bucket full of water at $20^\circ C$ will

- a) register a lower temperature than that of the water left in the bucket
b) have less thermal energy than that of the water left in the bucket
c) have less specific heat than that of the water left in the bucket

11) 243 K in degrees Celsius ($^\circ C$) is

$$T_C = T_K - 273 = 243 \text{ K} - 273 = \boxed{-30^\circ C}$$

12) The kinetic molecular theory can explain the existence of thermal energy in a substance.

- a) true** b) false

13) When using a thermometer for measuring the temperature of a substance, one should record the temperature

- a) about 2 minutes after the thermometer has been placed in the substance. That is usually the time needed for the liquid inside to expand or shrink completely.
b) until thermal equilibrium has been reached
c) until all the heat from the substance has been transferred to the thermometer
d) until the thermal energy from the substance is equal to the thermal energy of the thermometer
e) all of the answers

14) A calorimeter contains 1 kg of water at $20^\circ C$. A 0.20 kg block of zinc at $120^\circ C$ is placed in the water. What is the approximate final temperature of the system?

$$T_f = \frac{m_A c_A T_{Ai} + m_B c_B T_{Bi}}{m_A c_A + m_B c_B}$$

$$T_f = \frac{(1)(4180)(20^\circ C) + (0.2)(388)(120^\circ C)}{(1)(4180) + (0.2)(388)}$$

$$T_f = \frac{83600 + 9312}{4257.6} = \boxed{21.8^\circ C}$$

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web 1 packer brisket 12 to 14 pounds 3 cups jerk seasoning homemade or your favorite commercial brand 5 pimento leaves or bay leaves 2 tablespoons allspice berries preferably jamaican buttered toasted brioche rolls or

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web dec 19 2021 ingredients 1 medium onion finely chopped 1 2 cup scallions finely chopped 1 hot pepper finely chopped 3 tablespoons soy sauce 1 tablespoon oil 1 tablespoon cider vinegar or white vinegar 2 teaspoons fresh thyme leaves 2 teaspoons granulated sugar 1 teaspoon salt kosher salt or sea salt is

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web step 1 make the jerk seasoning place the chiles scallions garlic ginger cilantro thyme sugar salt allspice cinnamon pepper and salt in food processor and grind to a coarse paste work in the oil rum and soy sauce to obtain a thick but spreadable paste

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web recipe tips place the food in a glass baking dish and spread the jerk seasoning over it turning to ensure an even coating

marinate large pieces of meat overnight medium size pieces for 4 to 6 hours and small pieces for 1 to 2 hours covered in the refrigerator barbecue or indirect grill the meat using moderate heat and plenty of wood smoke

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web jan 14 2011 1 3 cup soy sauce 2 tablespoon vegetable oil 1 4 cup vinegar 1 onion 1 2 cup orange juice 2 cloves garlic 1 teaspoon grated ginger notes if you re using dried thyme use about 1 2 table spoon this makes about 3 cups of marinade and it can be stored in the fridge for about 1 month

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web in the process it s lost a lot of its fire spice and salt here s how they make it in the birthplace of jerk a town called boston beach on the north coast of jamaica count yourself lucky to have a blender in boston beach they grind the seasonings in a hand cranked spice mill

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web simply rub a small amount 1oz jerk seasoning 2 2lbs 1kg into your meat or fish for best results marinate overnight to absorb the spicy flavor and then cook until done on your barbeque or in your oven we are so proud that walkerswood jerk seasoning has been included in the list of 2015 men's health 125 best from walkerswood.com

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web for the brine 2 cups water 4 tablespoons coarse salt kosher or sea 4 tablespoons granulated sugar 6 garlic cloves peeled and halved lengthwise 1 small orange quartered 1 lemon quartered 1 small onion peeled and quartered 2

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