



# 11: THERMODYNAMICS

Physics Worksheet

Standard: 11th Science

Subject: Physics

Student Name: \_\_\_\_\_

**Q1.** Read the passage given below and answer the following questions from 1 to 5. **First Law of Thermodynamics** The first law of thermodynamics is the general law of conservation of energy applied to any system in which energy transfer from or to the surroundings (through heat and work) is taken into account. It states that the energy supplied to the system goes in partly to increase the internal energy of the system and the rest in work on the environment. Mathematically,  $\Delta Q = \Delta U + \Delta W$  where  $\Delta Q$  is the heat supplied to the system,  $\Delta W$  is the work done by the system and  $\Delta U$  is the change in internal energy of the system.  $\Delta Q$  and  $\Delta W$  depend on the path taken to go from initial to final states, but the combination  $\Delta Q - \Delta W$  is path independent.

1. The first law of thermodynamics is concerned with conservation of:

1. Number of molecules
2. Number of moles
3. Energy
4. Temperature

2. Which of the following is not a path function?

1.  $\Delta Q$
2.  $\Delta Q + \Delta W$
3.  $\Delta W$
4.  $\Delta Q - \Delta W$

3. An electric heater supplies heat to a system at a rate of  $120W$ . If system performs work at a rate of  $80 \text{ J s}^{-1}$ , the rate of increase in internal energy is:

1.  $30 \text{ J s}^{-1}$
2.  $40 \text{ J s}^{-1}$
3.  $50 \text{ J s}^{-1}$
4.  $60 \text{ J s}^{-1}$

4. A system goes from A to B by two different paths in the  $P - V$  diagram as shown in figure. Heat given to the system in path 1 is  $1100 \text{ J}$ , the work done by the system along path 1 is more than path 2 by  $150 \text{ J}$ . The heat exchanged by the system in path 2 is:



# Physics Thermodynamics Test Answers

**Yoni Kahn, Adam Anderson**



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### **Table of Contents Physics Thermodynamics Test Answers**

1. Understanding the eBook Physics Thermodynamics Test Answers
  - The Rise of Digital Reading Physics Thermodynamics Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Physics Thermodynamics Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physics Thermodynamics Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physics Thermodynamics Test Answers
  - Personalized Recommendations
  - Physics Thermodynamics Test Answers User Reviews and Ratings
  - Physics Thermodynamics Test Answers and Bestseller Lists

5. Accessing Physics Thermodynamics Test Answers Free and Paid eBooks
  - Physics Thermodynamics Test Answers Public Domain eBooks
  - Physics Thermodynamics Test Answers eBook Subscription Services
  - Physics Thermodynamics Test Answers Budget-Friendly Options
6. Navigating Physics Thermodynamics Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Physics Thermodynamics Test Answers Compatibility with Devices
  - Physics Thermodynamics Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physics Thermodynamics Test Answers
  - Highlighting and Note-Taking Physics Thermodynamics Test Answers
  - Interactive Elements Physics Thermodynamics Test Answers
8. Staying Engaged with Physics Thermodynamics Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physics Thermodynamics Test Answers
9. Balancing eBooks and Physical Books Physics Thermodynamics Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physics Thermodynamics Test Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physics Thermodynamics Test Answers
  - Setting Reading Goals Physics Thermodynamics Test Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physics Thermodynamics Test Answers
  - Fact-Checking eBook Content of Physics Thermodynamics Test Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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