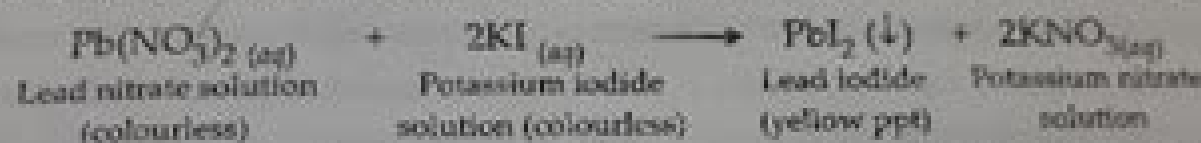


• Precipitation Reactions

When the aqueous solution of two ionic compounds react by exchanging their ions/radicals, to form two or more new compounds, such that one of the products formed is an insoluble salt, and hence, forms precipitate, the double displacement reaction is said to be **precipitation reaction**.

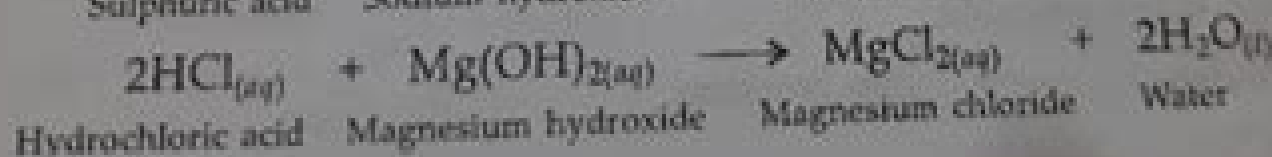
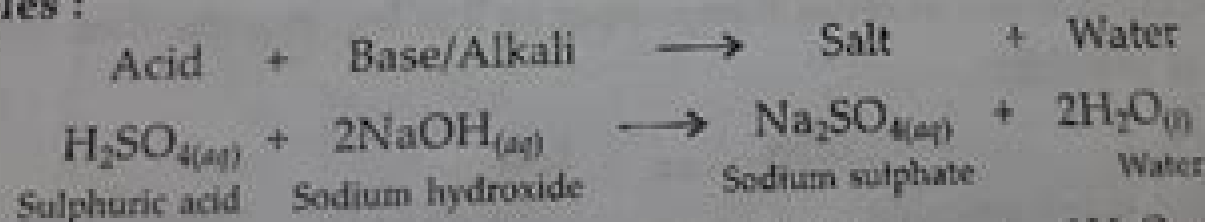
For example : When lead nitrate solution is mixed with potassium iodide solution a yellow precipitate is formed. This reaction is a precipitation reaction and can be expressed as follows:



In the above reaction, the lead nitrate solution and potassium iodide solution exchange their ions to form a precipitate of lead iodide and potassium nitrate solution.

When an aqueous solution of an acid reacts with a base (alkali) by exchanging their ions/radicals to form salt and water as the only products, the reaction which takes place is called **neutralisation reactions**.

Examples :



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