

## Chapter 10 - Chemical Quantities

### 10.1 The Mole: A Measurement of Matter:

Mole: (Frank Green's videos)

- Unit of measurement

**Representative Particles:**

- Smallest unit of a substance (atom, molecule, formula unit, ions, anything...)
- Avogadro's number =  $6.02 \times 10^{23}$  representative particles (RP)
- $6.02 \times 10^{23}$  representative particles = 1 mol

**Molar Mass:**

- Molar mass = g/mol
- Molar mass of:
  - an element = atomic mass (g)  
Ex. 24.3 g/mol = molar mass of Mg
  - a compound = sum of the atomic masses (g) of each element  
Ex. 58.44g/mol = molar mass of NaCl
- Molar mass = 1 mol

**Molar Volume:**

- Molar volume = 22.4 L/mol
- Standard temperature and pressure (STP) =  
temp =  $0^{\circ}\text{C}$   
pressure = 1 atm
- 22.4 L = 1 mol (at STP)

**Calculating Molar mass from Density:**

- molar mass (g/mole) = density at STP (g/L)  $\times$  molar volume at STP (22.4L/mole)

$$\text{Molar mass (g/mol)} = \left( \frac{\text{grams}}{\text{L}} \right) \left( \frac{22.4 \text{ L}}{\text{mole}} \right)$$

**Calculating Density from Molar mass:**

- Density at STP (g/L) = molar mass (g/mole)  $\times$  molar volume at STP (1mole/22.4L)

$$\text{Density (g/L)} = \left( \frac{\text{grams}}{1 \text{ mole}} \right) \left( \frac{1 \text{ mole}}{22.4 \text{ L}} \right)$$

# Prentice Hall Chemistry Chemical Quantities Chapter 10

## Lecture Notes

**Mark Steyvers**



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