

Name: _____

Date: _____

Period: _____

pH and pOH Calculations

Find the pH and the pOH of the solution	pH	pOH
1. A 0.023 M solution of perchloric acid (HClO_4)	$\text{pH} = -\log(0.023)$ $\text{pH} = 1.638 \rightarrow 1.6$	$\text{pOH} = 14 - 1.6$ $\text{pOH} = 12.4$
2. A 6.6×10^{-6} M solution of nitric acid (HNO_3)	$\text{pH} = -\log(6.6 \times 10^{-6})$ $\text{pH} = 5.18 \rightarrow 5.2$	$\text{pOH} = 14 - 5.2$ $\text{pOH} = 8.8$
3. A 0.0334 M solution of potassium hydroxide (KOH)	$\text{pH} = 14 - 1.48$ $\text{pH} = 12.52 \rightarrow 12.5$	$\text{pOH} = -\log(0.0334)$ $\text{pOH} = 1.476 \rightarrow 1.48$
4. A 1.0×10^{-5} M solution of hydroarsenic acid (H_3As)	$\text{pH} = -\log(3.0 \times 10^{-5})$ $\text{pH} = 4.52 \rightarrow 4.5$	$\text{pOH} = 14 - 4.5$ $\text{pOH} = 9.5$
5. A 2.23×10^{-12} M solution of aluminum hydroxide $\text{Al}(\text{OH})_3$	$\text{pH} = 14 - 1.17$ $\text{pH} = 12.83 \rightarrow 12.8$	$\text{pOH} = -\log(6.69 \times 10^{-12})$ $\text{pOH} = 1.17$
6. A 0.000901 M solution of magnesium hydroxide $\text{Mg}(\text{OH})_2$	$\text{pH} = 14 - 2.74$ $\text{pH} = 11.26 \rightarrow 11.3$	$\text{pOH} = -\log(0.001802)$ $\text{pOH} = 2.74$
7. A 7.23×10^{-5} M solution of thiosulfic acid ($\text{H}_2\text{S}_2\text{O}_3$)	$\text{pH} = -\log(14.46 \times 10^{-5})$ $\text{pH} = 3.839 \rightarrow 3.84$	$\text{pOH} = 14 - 3.84$ $\text{pOH} = 10.16 \rightarrow 10.2$
8. A 2.58×10^{-5} M solution of oxalic acid ($\text{H}_2\text{C}_2\text{O}_4$)	$\text{pH} = -\log(5.16 \times 10^{-5})$ $\text{pH} = 5.287 \rightarrow 5.29$	$\text{pOH} = 14 - 5.29$ $\text{pOH} = 8.71$
9. A 0.000469 M solution of sodium hydroxide (NaOH)	$\text{pH} = 14 - 3.33$ $\text{pH} = 10.67 \rightarrow 10.7$	$\text{pOH} = -\log(0.000469)$ $\text{pOH} = 3.328 \rightarrow 3.33$
10. A 0.000896 M solution of sulfuric acid (H_2SO_4)	$\text{pH} = -\log(0.01792)$ $\text{pH} = 2.746 \rightarrow 2.75$	$\text{pOH} = 14 - 2.75$ $\text{pOH} = 11.25 \rightarrow 11.3$

Ph And Poh Calculation Answer Key

**John Olmsted, Greg Williams, Robert C.
Burk**



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Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Chemistry and Physics for Nurse Anesthesia, Second Edition* David Shubert, PhD, John Leyba, PhD, 2013-03-15 Print CourseSmart

Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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