

## QUESTION 8

- 8.1 8,78 g of oxygen gas is required to react with an unknown mass of hydrogen gas to form water vapour. The balanced equation for this is:



- 8.1.1 Define the term *exothermic reaction*. (2)
- 8.1.2 Is this reaction a synthesis or a decomposition reaction? Give a reason for the answer. (2)
- 8.1.3 Calculate the number of moles of water vapour formed. (2)
- 8.2 The reaction between zinc and dilute hydrochloric acid is represented by the balanced equation below:



The concentration of hydrochloric acid solution is  $0,1 \text{ mol.dm}^{-3}$ . The mass of zinc chloride formed is 0,85 g.

- 8.2.1 Define the term *concentration* in words. (2)
- 8.2.2 Is the above reaction an example of a redox reaction? Write down **YES** or **NO** and give a reason for the answer. (2)
- 8.2.3 Calculate the volume of hydrochloric acid solution needed to react with the zinc to form 0,85 g of  $ZnCl_2$ . (5)

[15]

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