

Learning Guide in Modern Physics

Photoelectric Effect

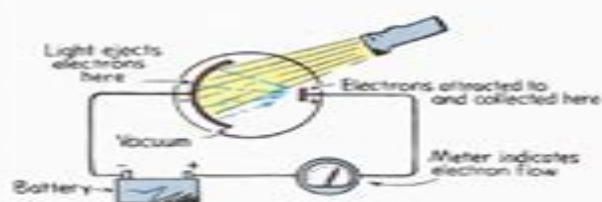


Figure 31.1 An apparatus used for observing the photoelectric effect. Reversing the polarity and stopping the electron flow provides a way to measure the energy of the electrons.

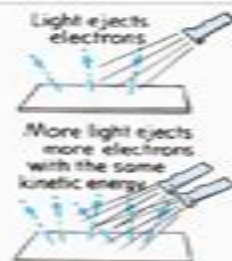


FIGURE 31.2
The photoelectric effect depends on intensity.



FIGURE 31.3
The photoelectric effect depends on frequency.

CHECK YOURSELF

1. Will brighter light eject more electrons from a photosensitive surface than dimmer light of the same frequency?
2. Will high-frequency light eject a greater number of electrons than low-frequency light?

CHECK YOUR ANSWERS

1. Yes. The number of ejected electrons depends on the number of incident photons.
2. Not necessarily. The energy (not the number) of ejected electrons depends on the frequency of the illuminating photons. A bright source of blue light, for example, may eject more electrons at a lower energy than a dim violet source.

In the latter part of the nineteenth century, several investigators noticed that light was capable of ejecting electrons from various metal surfaces. This is the photoelectric effect, now used in electric eyes, in the photographer's light meter, and in "reading" the sound tracks of motion pictures.

An arrangement for observing the photoelectric effect is shown in Figure 31.1. Light shining on the negatively charged, photosensitive metal surface liberates electrons. The liberated electrons are attracted to the positive plate and produce a measurable current. If we instead charge this plate with just enough negative charge that it repels electrons, the current can be stopped. We can then calculate the energies of the ejected electrons from the easily measured potential difference between the plates.

The photoelectric effect was not particularly surprising to early investigators. The ejection of electrons could be accounted for by classical physics, which pictures the incident light waves building an electron's vibration up to greater and greater amplitudes until it finally breaks loose from the metal surface, just as water molecules break loose from the surface of heated water. It should take considerable time for a weak source of light to give electrons in a metal enough energy to make them boil off the surface. Instead, it was found that electrons are ejected as soon as the light is turned on—but not as many as are ejected as with a strong light source. Careful examination of the photoelectric effect led to several observations that were quite contrary to the classical wave picture:

1. The time lag between turning on the light and the ejection of the first electrons was unaffected by the brightness or frequency of the light.
2. The effect was easy to observe with violet or ultraviolet light but not with red light.
3. The rate at which electrons were ejected was proportional to the brightness of the light.
4. The maximum energy of the ejected electrons was unaffected by the brightness of the light. However, there were indications that the electron's energy did depend on the frequency of the light.

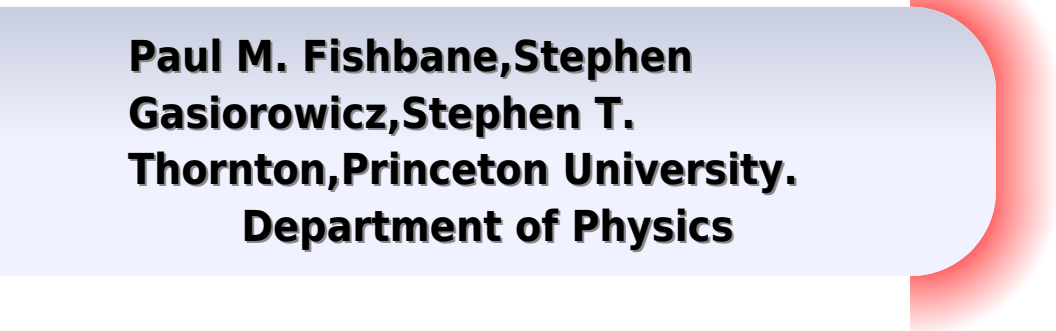
The lack of any appreciable time lag was especially difficult to understand in terms of the wave picture. According to the wave theory, an electron in dim light should, after some delay, accumulate sufficient vibrational energy to fly out, while an electron in bright light should be ejected almost immediately. However, this didn't occur. It was not unusual to observe an electron being ejected immediately, even under the dimmest light. The observation that the brightness of light in no way affected the energies of ejected electrons was also perplexing. The stronger electric fields of brighter light did not cause electrons to be ejected at greater speeds. More electrons were ejected in brighter light, but not at greater speeds. A weak beam of ultraviolet light, on the other hand, produced a smaller number of ejected electrons but at much higher speeds. This was most puzzling.

Einstein produced the answer in 1905, the same year he explained Brownian motion and set forth his theory of special relativity. His clue was Planck's quantum theory of radiation. Planck had assumed that the emission of light in quanta was due to restrictions on the vibrating atoms that produced the light. That is, he assumed that energy in matter is quantized, but that radiant energy is continuous. Einstein, on the other hand, attributed quantum properties to light itself and viewed radiation as a hail of particles. To emphasize this particle aspect, we speak of photons (by analogy with electrons, protons, and neutrons) whenever we are thinking of the particle nature of light. One photon is completely absorbed by each electron ejected from the metal. The absorption is an all-or-nothing process and is immediate, so there is no delay as "wave energies" build up.

The photoelectric effect proves conclusively that light has particle properties. We cannot conceive of the photoelectric effect on the basis of waves. On the other hand, we have seen that the phenomenon of interference demonstrates convincingly that light has wave properties. We cannot conceive of interference in terms of particles. In classical physics, this appears to be and is contradictory. From the point of view of quantum physics, light has properties resembling both. It is "just like a wave" or "just like a particle," depending on the particular experiment. So we think of light as both, as a wave-particle. How about "wavicle"? Quantum physics calls for a new way of thinking.

Physics Learning Guide Answers

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