

Name: \_\_\_\_\_

## Chemistry Practice: Balancing Chemical Equations

1. Balance the equations below. Remember this means only using coefficients. All formulas are written correctly.

- a.)  $2 \text{NaHCO}_3 \longrightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- b.)  $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$
- c.)  $2 \text{H}_2 + \text{O}_2 \longrightarrow 2 \text{H}_2\text{O}$
- d.)  $2 \text{FeO} + \text{C} \longrightarrow 2 \text{Fe} + \text{CO}_2$
- e.)  $\text{Cl}_2 + 2 \text{KI} \longrightarrow 2 \text{KCl} + \text{I}_2$
- f.)  $\text{B}_2\text{O}_3 + 3 \text{Mg} \longrightarrow 3 \text{MgO} + 2 \text{B}$
- g.)  $\text{Ca} + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + 2 \text{H}_2$
- h.)  $4 \text{H}_2\text{O} + \text{Na}_2\text{B}_4\text{O}_7 + \text{H}_2\text{SO}_4 \longrightarrow 4 \text{H}_3\text{BO}_3 + \text{Na}_2\text{SO}_4$
- i.)  $\text{CaC}_2 + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- j.)  $\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$
- k.)  $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{SO}_2 + \text{H}_2\text{O}$
- l.)  $\text{H}_2\text{SO}_4 + 2 \text{NaCl} \longrightarrow 2 \text{HCl} + \text{Na}_2\text{SO}_4$
- m.)  $2 \text{NaCl} + 2 \text{H}_2\text{O} \longrightarrow 2 \text{NaOH} + \text{Cl}_2 + \text{H}_2$
- n.)  $3 \text{BaO} + 2 \text{Al} \longrightarrow 3 \text{Ba} + \text{Al}_2\text{O}_3$
- o.)  $3 \text{O}_2 \longrightarrow 2 \text{O}_3$
- p.)  $\text{SiO}_2 + 2 \text{C} \longrightarrow 2 \text{CO} + \text{Si}$
- q.)  $2 \text{PbS} + 3 \text{O}_2 \longrightarrow 2 \text{PbO} + 2 \text{SO}_2$
- r.)  $2 \text{HgO} \longrightarrow 2 \text{Hg} + \text{O}_2$
- s.)  $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12 \text{O}_2 \longrightarrow 12 \text{CO}_2 + 11 \text{H}_2\text{O}$
- t.)  $\text{C}_3\text{H}_8 + 5 \text{O}_2 \longrightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- u.)  $2 \text{FeCl}_3 + 3 (\text{NH}_4)_2\text{S} \longrightarrow 6 \text{NH}_4\text{Cl} + \text{Fe}_2\text{S}_3$
- v.)  $\text{P}_2\text{O}_5 + 6 \text{NaOH} \longrightarrow 2 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
- w.)  $3 \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 \longrightarrow 3 \text{BaSO}_4 + 2 \text{AlCl}_3$
- x.)  $4 \text{NH}_3 + 5 \text{O}_2 \longrightarrow 6 \text{H}_2\text{O} + 4 \text{NO}$
- y.)  $2 \text{Na}_2\text{O}_2 + 2 \text{H}_2\text{O} \longrightarrow 4 \text{NaOH} + \text{O}_2$
- z.)  $\text{Sn(ClO}_3)_4 \longrightarrow \text{SnCl}_4 + 6 \text{O}_2$

Text Problems to be Completed on the Other Side of this Worksheet:

\*P. 173 #45,49,53,54,56,57,58,63, 65,66,67,71

\*(Although the directions may say to write an "unbalanced" equation, write a balanced equation for each.)

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**Dr. Sandeep Ahlawat**



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