

## Worksheet #10: Stoichiometry Supplement

### Limiting Reactant, Theoretical Yield, and Percent Yield Step-by-Step

Consider the reaction between magnesium and nitrogen, which produces magnesium nitride:



Molar Masses needed

	Mg	N <sub>2</sub>	Mg <sub>3</sub> N <sub>2</sub>
g/mol	24.31	28.02	100.95
moles			
grams	14.32		

Track calculations as  
needed (if you want)

Problem A: Calculate  
Reactant Needed and

# Supplemental Stoichiometry Problems

**SA Adler**



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