

Name: \_\_\_\_\_

## Neutralization Reactions Worksheet

Reading: Cycle 6 Chemistry II Lesson 4

Neutralization Reaction: Acid + Base  $\rightarrow$  Salt + Water

H<sup>+</sup> from acid and OH<sup>-</sup> from base combine to form H<sub>2</sub>O. Leftovers form 'salt'.

Remember in salt, write metal first! KCl not ClK!

**Example 1:**  $\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

**Example 2:**  $\_\_\text{HF} + \_\_\text{Ca(OH)}_2 \rightarrow \_\_\_\_\_\_ + \_\_\_\_\_\_$

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)<sub>2</sub>

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow \_\_\_\_\_\_ + \_\_\_\_\_\_$

Two H and two OH make two waters. Think of prom! Two boys and two girls make two couples.

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow \underline{2\text{ H}_2\text{O}} + \_\_\_\_\_\_$

What's left? 2 F atoms and 1 Ca atom form the 'salt'. Remember that Ca is the metal, so it goes first.

$\underline{2}\text{ HF} + \underline{1}\text{ Ca(OH)}_2 \rightarrow \underline{2\text{ H}_2\text{O}} + \underline{\text{CaF}_2}$

Safety check: is CaF<sub>2</sub> a good formula? P.T. tells you ions are Ca<sup>+2</sup> and F<sup>-</sup>, so total charge is 0 like it should be.

1. Complete & balance these neutralization reactions. Remember that H<sub>2</sub>O = HOH!

$\text{H}_2\text{SO}_4 + \text{Sr(OH)}_2 \rightarrow \_\_\text{H}_2\text{O} + \text{SrSO}_4$

$\_\_\text{Ca(OH)}_2 + \_\_\text{H}_3\text{PO}_4 \rightarrow \_\_\text{H}_2\text{O} + \_\_\text{Ca}_3(\text{PO}_4)_2$

$\_\_\text{HBr} + \text{Ba(OH)}_2 \rightarrow \_\_\text{H}_2\text{O} + \_\_\_\_\_\_$

$\text{Zn(OH)}_2 + \_\_\text{HNO}_3 \rightarrow \_\_\_\_\_\_ + \text{Zn(NO}_3)_2$

$\text{Al(OH)}_3 + \_\_\text{HCl} \rightarrow \_\_\_\_\_\_ + \_\_\_\_\_\_$

2. Complete and balance the following equations representing neutralization reactions:

a)  $\_\_\text{CsOH} + \text{H}_2\text{SO}_3 \rightarrow \_\_\text{H}_2\text{O} + \text{Cs}_2\text{CO}_3$

b)  $\_\_\text{HF} + \text{Mg(OH)}_2 \rightarrow \_\_\text{H}_2\text{O} + \_\_\_\_\_\_$

c)  $\_\_\text{HNO}_3 + \text{Al(OH)}_3 \rightarrow \_\_\text{H}_2\text{O} + \text{Al(NO}_3)_3$

d)  $\text{HCl} + \text{KOH} \rightarrow \_\_\_\_\_\_ + \_\_\_\_\_\_$

e)  $\text{HBrO}_3 + \text{LiOH} \rightarrow \text{H}_2\text{O} + \_\_\_\_\_\_$

# Section 19 4 Neutralization Reactions Worksheet

## Answers

**Jianjun Gao**



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