

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$   
 $V_2 = 1.47 \text{ L}; P_2 = ? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1 V_1 = P_2 V_2$   
Isolate for  $P_2$ :  $P_2 = (P_1 V_1)/V_2$   
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$   
 $V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$   
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$   
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$   
Get rid of P  
 $V_1/T_1 = V_2/T_2$   
Isolate  $V_2$ :  $V_2 = (V_1/T_1)T_2$   
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV = nRT$  will work. ( $n$  = moles and  $R$  = 0.08206 L atm/K mol)  
 $P = 1.04 \text{ atm}, V = ? \text{ L}, n = \text{OOPS we need to convert from grams to moles!}, T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$   
 $n$ : Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$   
Isolate  $V$ :  $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

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**L Reisser**



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