

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

1. Explain how a chemical equation is written.
2. Explain how to write a balanced equation.
3. Write the steps for writing a balanced chemical equation.

Vocabulary

- | | |
|----------------------|----------------------|
| 1. chemical equation | 4. coefficient |
| 2. stoichiometric | 5. balanced equation |
| 3. reactant | |

Part A Completion

Use the vocabulary words in each sentence, writing either the correct word form or the correct chemical equation. Write each answer on a separate sheet of paper and underline the answer.

1. A chemical equation that includes information about the _____ of the reactants, the products, and the states of matter is called a _____ equation. The word coefficient refers to the whole number in the _____ of a chemical equation. The law of conservation of _____ is demonstrated by the fact that the number of _____ atoms on the left side of a balanced equation is equal to the number of _____ atoms on the right side of the equation. A chemical equation is balanced when the number of _____ of each element is the same on both sides of the equation.

2. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____, where the coefficient is _____, the value _____, and the gas _____ is a diatomic molecule. The reaction is balanced when the number of _____ of each element is the same on both sides of the equation. The balanced equation is _____.

3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____

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Michael J. Padilla



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