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VS800 / Boulevard / S50
2005 - 2009

Suzuki Boulevard S50 User Manual

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monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2
was produced i what mass of iodine was produced

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160.0 \text{ g } \text{Al} \times \frac{1 \text{ mol } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 150.1 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\frac{126.0 \text{ g } \text{Al}_2\text{O}_3}{150.1 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 83.9\%$

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