

## Chapter 19 Study Guide: Acids, Bases &amp; Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
Brønsted-Lowry	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

- a.  $\text{H}_2\text{S}$  – hydrosulfuric acid  
 b.  $\text{HC}_2\text{H}_3\text{O}_2$  – acetic acid  
 c.  $\text{NaOH}$  – sodium hydroxide  
 d.  $\text{HNO}_3$  – nitric acid  
 e. Nitrous acid –  $\text{HNO}_2$   
 f. Calcium hydroxide –  $\text{Ca}(\text{OH})_2$   
 g. hydrochloric acid –  $\text{HCl}$

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$M_A \cdot V_A \cdot n_A = M_B \cdot V_B \cdot n_B$$

4. What is the pH of
- $1.34 \times 10^{-4}$
- M solution of hydrochloric acid?

$$\text{pH} = -\log [1.34 \times 10^{-4} \text{M}] = 3.873$$

5. Determine the pH and pOH of a
- $1.85 \times 10^{-6}$
- M NaOH solution.

$$\text{pOH} = -\log [1.85 \times 10^{-6} \text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the
- $[\text{OH}^-]$
- of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9} \text{ M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35 \text{g NaOH}}{2.000 \text{L}} \times \frac{1 \text{mol NaOH}}{40.00 \text{g NaOH}} = 0.44 \text{ M NaOH}$$

$$\text{pOH} = -\log[0.44 \text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Study Guide Acids Bases And Titration

**JS Bruner**



## **Study Guide Acids Bases And Titration:**

## **Study Guide Acids Bases And Titration** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Study Guide Acids Bases And Titration**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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